

Thermochemistry Practice Worksheet

1. How much heat is released during the combustion of 1.2 g of hexane if the molar heat of reaction is 1510 kJ/mol.

a. Is the equation exothermic or endothermic?

b. Which of the potential energy diagrams below represents the reaction?



2. Which type of enthalpy reactions are *always* endothermic?
3. Which of the following represents a ΔH_{vap}
- Freezing of water
 - Addition of water
 - Combustion of ethane
 - Production of steam from water
4. If 0.315 moles of hexane (C_6H_{14}) undergoes combustion in a bomb calorimeter containing 5.65 liters of water, calculate the molar heat of combustion of hexane if the water temperature rises 55.4 °C? The specific heat capacity of water is 4.184 J/g °C.
5. The mixture between potassium perchlorate and aluminum is the standard today in the fireworks industry since it is very powerful and safer than most other mixtures. The reaction between these species takes the following form:



Use enthalpies of formation shown below to determine the ΔH for this reaction.

KClO_4 : -436.5 kJ/mol KCl : -432.8 kJ/mol Al_2O_3 : -1675.7 kJ/mol

6. Self-heating cans may be used to warm drinks such as coffee. When the button on the can is pushed, a seal is broken, allowing water and calcium oxide to mix and react. The reaction produces solid calcium hydroxide and releases heat. If more water is used the calcium hydroxide is produced as a solution instead of as a solid.



Using the following data, calculate the enthalpy change, in kJ mol⁻¹, for this reaction.



7. Sketch the changes in a Maxwell-Boltzmann Distribution graph at different temperatures.
8. Using the values given to construct a Born-Haber cycle and calculate the ΔH_f° for MgCl_2 .

$\Delta H_{\text{atom}} = 150 \text{ kJ/mol}$ | $\Delta H_{\text{IE1}} = 736 \text{ kJ/mol}$ | $\Delta H_{\text{IE2}} = 1450 \text{ kJ/mol}$ | $\Delta H_{\text{E}} = -364 \text{ kJ/mol}$

$\Delta H_{\text{lat}} = 2493 \text{ kJ/mol}$ | $E = 242 \text{ kJ/mol}$

9. Which ionic compound has the largest amount of lattice energy?
- NaF
 - AlF_3
 - AlN
 - MgF_2
10. Calculate the enthalpy change from bond energies for each of these reactions:
- $\text{H}_{2(\text{g})} + \text{F}_{2(\text{g})} \rightarrow 2 \text{HF}_{(\text{g})}$

Thermochemistry Chemical Reaction Review

J Ma

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Thermochemistry Chemical Reaction Review:

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