

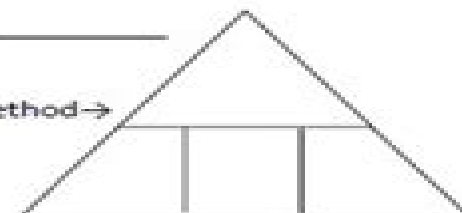
Thermochemistry Unit Review Packet

Name: _____

Date: _____

Part 1 Specific Heat Equation Problems

Set up and use the triangle method →



The specific heat equation is $Q = m \cdot c \cdot \Delta T$.

Q = heat energy, m = mass, and ΔT = change in temperature ($T_{\text{final}} - T_{\text{initial}}$)

Show your work and label all of the units of measurement throughout.

1. How many joules of heat are needed to raise the temperature of 52.0 g of zirconium from 25°C to 56°C, if the specific heat of zirconium is 0.27 J/g°C?
2. A 3.65 g piece of iron absorbs 144.7 joules of heat energy and its temperature changes from 23.0°C to 111°C. Calculate the specific heat capacity of iron.

Part 2 Gibbs Free Energy Equation Problems

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta H = T\Delta S$$

$$\Delta S = \frac{\Delta H}{T}$$

$$^{\circ}\text{C} + 273.2 = ^{\circ}\text{K}$$

1. For the reaction $\text{NH}_4\text{Br}(s) \rightarrow \text{NH}_3(g) + \text{HBr}(g)$, at 302.7°K, the enthalpy change is 158 kJ/mol, and the change in entropy is 0.422 kJ/mol·°K. Calculate ΔG .
2. Calculate the entropy of the vaporization phase change for silver at 2453°K. The change in enthalpy is +250.58 kJ/mol.

Thermochemistry Review Packet

J Dewey



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Table of Contents Thermochemistry Review Packet

1. Understanding the eBook Thermochemistry Review Packet
 - The Rise of Digital Reading Thermochemistry Review Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Review Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Review Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Review Packet
 - Personalized Recommendations
 - Thermochemistry Review Packet User Reviews and Ratings
 - Thermochemistry Review Packet and Bestseller Lists
5. Accessing Thermochemistry Review Packet Free and Paid eBooks
 - Thermochemistry Review Packet Public Domain eBooks
 - Thermochemistry Review Packet eBook Subscription Services
 - Thermochemistry Review Packet Budget-Friendly Options

6. Navigating Thermochemistry Review Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Review Packet Compatibility with Devices
 - Thermochemistry Review Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Review Packet
 - Highlighting and Note-Taking Thermochemistry Review Packet
 - Interactive Elements Thermochemistry Review Packet
8. Staying Engaged with Thermochemistry Review Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Review Packet
9. Balancing eBooks and Physical Books Thermochemistry Review Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Review Packet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Review Packet
 - Setting Reading Goals Thermochemistry Review Packet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Review Packet
 - Fact-Checking eBook Content of Thermochemistry Review Packet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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