- The enthalpy of formation for silver nitrate is -2304 kJ/mol.
 - a. Write the balanced reaction representing this enthalpy.
 - b. What is the value of the enthalpy to form three (3) moles of silver nitrate?
 - c. What is the value of the enthalpy to form 15.490 grams of silver nitrate?

Amswers

- a. $Ag(s) + \frac{1}{2} N_2(g) + \frac{3}{2} O_2(g) \rightarrow 1$ $AgNO_3(s)$; $\Delta H = -2304 \text{ kJ/mol}$
- b. (3 mol)(-2304 kJ/mol) = -6912 kJ
- c. $\frac{15.490g}{169.88g} \left(\frac{1mol}{1mol} \right) \left(\frac{-2304kJ}{1mol} \right) = -210.08kJ$
- 2. When 1.00 L of 1.00 M Ba(NO₃)₂ solution at 25.0°C is mixed with 1.00 L of 1.00 M Na₂SO₄ solution at 25.0°C in a calorimeter, the white solid BaSO₄ forms and the temperature of the mixture increases to 28.1°C. Assuming that the calorimeter absorbs only a negligible quantity of heat, that the specific heat capacity of the solution is 4.18J/°C•g, and that the density of the final mixture is 1.0 g/mL, calculate the enthalpy change per mole of NaNO₃ formed.

Almswers

 $\frac{1}{Q}$ Ba(NO₃)₂ + $\frac{1}{2}$ Na₂SO₄ → $\frac{2}{2}$ NaNO₃ + $\frac{1}{2}$ BaSO₄ $\frac{1}{Q}$ Ba(NO₃)₂ + $\frac{1}{2}$ Na₂SO₄ → $\frac{1}{2}$ NaNO₃ + $\frac{1}{2}$ BaSO₄ $\frac{1}{2}$ BaSO₄ $\frac{1}{2}$ BaSO₄ $\frac{1}{2}$ BaSO₄ $\frac{1}{2}$ Select one of the reactants to obtain moles: $\frac{1}{2}$ (1.00 L)(1.00 M) = 1.00 mol Ba(NO₃)₂ $\frac{1}{2}$ (1.00 L)(1.00 M) = 1.00 mol Ba(NO₃)₂

 $\left(\frac{+25916J}{1mol \operatorname{Ba}(NO_3)_2}\right)\left(\frac{1mol \operatorname{Ba}(NO_3)_2}{2mol \operatorname{Ba}(NO_3)_2}\right) = +12958J$ $\left[1.30 \times 10^4 \operatorname{J}\right]$

The molar heat capacity of mercury is 28.1 J/mol·K. What is the specific heat capacity of this metal?

dissector

3.

- Specific heat capacity has units of J/g. C.
- Since the temperature is a **change in temperature**, rather than a direct temperature, the unit is based on the difference between two temperatures. Since "C and K are converted through addition or subtraction, the difference between two temperatures in "C is the same difference between the same temperatures in K. Take for instance T_i = 25 "C and T_f = 55 "C. The difference between these two is T_f T_i = 55 25 = 30 "C. At the same time, these temperatures in Kelvin would appear as T_i = 298K and T_f = 328K. The difference between these two is T_f T_i = 328 298 = 30K. The magnitude is the same. Therefore, the molar heat capacity of mercury can also be represented as J/mol-"C.

Now, the moles can be converted to grams:

 $\left(\frac{28.1J}{mol \cdot C}\right)\left(\frac{1mol \operatorname{Hg}}{200.59g \operatorname{Hg}}\right) = 0.140 \frac{J}{g^{-1}C}$

Melonie Anne Teichert

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