

Thermochemistry Worksheet Key

1. How much heat, in calories, is released or absorbed if a 15.0 g sample of copper is converted from
(a) the solid to liquid state.

$$q = m \times \Delta H_{fus} = 15.0 \text{ g} \times 206 \frac{\text{J}}{\text{g}} \times \frac{1 \text{ cal}}{4.184 \text{ J}} = 738.5 = 739 \text{ cal}$$

So 739 cal is absorbed

- (b) the gas to the liquid state.

$$\begin{aligned} q &= m \times \Delta H_{con} = 15.0 \text{ g} \times -4.7 \frac{\text{kJ}}{\text{g}} \times \frac{1000 \text{ J}}{1 \text{ kJ}} \times \frac{1 \text{ cal}}{4.184 \text{ J}} = -16,849 \\ &= -1.68 \times 10^4 \text{ cal or } -16.8 \text{ kcal} \end{aligned}$$

So 16.8 kcal is released.

For copper m.p. = 1084 °C ; b.p. = 2927 °C; ΔH_{fus} = 206 J/g, ΔH_{vap} = 4.7 kJ/g

2. How much heat is absorbed or released if a 12.5 g sample of aluminum changes temperature from 25.0 to 87.5 °C? The specific heat capacity of aluminum is 0.900 J/g K

$$T_{final} = 87.5 + 273.15 = 360.65 \text{ K} \quad T_{initial} = 25.0 + 273.15 = 298.15 \text{ K}$$

$$\begin{aligned} q &= m \times c \times \Delta T = 12.5 \text{ g} \times 0.900 \frac{\text{J}}{\text{g} \cdot \text{K}} \times (360.65 - 298.15) \text{ K} \\ &= 12.5 \text{ g} \times 0.900 \frac{\text{J}}{\text{g} \cdot \text{K}} \times 62.5 \text{ K} = 703.1 = 703 \text{ J} \end{aligned}$$

3. What was the mass of a sample of water if the addition of 4.28×10^3 joules raised the temperature of the sample from 22.0 °C to 34.5 °C?

$$T_{final} = 34.5 + 273.15 = 307.65 \text{ K} \quad T_{initial} = 22.0 + 273.15 = 295.15 \text{ K}$$

T increases so endothermic so $q > 0$

$$q = m \times c \times \Delta T$$

$$\text{so } m = \frac{q}{c \times \Delta T} = \frac{4.28 \times 10^3 \text{ J}}{4.184 \frac{\text{J}}{\text{g} \cdot \text{K}} \times (307.65 - 295.15) \text{ K}} = \frac{4.28 \times 10^3 \text{ J}}{4.184 \frac{\text{J}}{\text{g} \cdot \text{K}} \times 12.5 \text{ K}} = 81.83 = 81.8 \text{ g}$$

Thermochemistry Practice Problems Worksheet

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Thermochemistry Practice Problems Worksheet:

Unveiling the Magic of Words: A Report on "**Thermochemistry Practice Problems Worksheet**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Thermochemistry Practice Problems Worksheet**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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