

Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone, CaCO_3 , to lime, CaO , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{CaCO}_3(\text{s}) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CaO}(\text{s}) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CO}_2(\text{g}) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4. $\text{CH}_4(\text{g}) + 3 \text{Cl}_2(\text{g}) \rightarrow \text{CHCl}_3(\text{g}) + 3 \text{HCl}(\text{g}) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide, CH_2N_2 , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: $\Delta H^\circ_f \text{NO}(\text{g}) = 90.3 \text{ kJ/mol}$ and $\Delta H^\circ_f \text{H}_2\text{O}(\text{g}) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction: $2 \text{Al}(\text{s}) + \text{Fe}_2\text{O}_3(\text{s}) \rightarrow \text{Al}_2\text{O}_3(\text{s}) + 2 \text{Fe}(\text{s}) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe_2O_3 , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

Thermochemistry Practice Calculation Unit 1

Yijin Wang



Thermochemistry Practice Calculation Unit 1:

Embracing the Track of Phrase: An Emotional Symphony within **Thermochemistry Practice Calculation Unit 1**

In a global taken by monitors and the ceaseless chatter of instant interaction, the melodic beauty and psychological symphony produced by the published word usually diminish into the backdrop, eclipsed by the persistent sound and distractions that permeate our lives. But, located within the pages of **Thermochemistry Practice Calculation Unit 1** an enchanting literary prize brimming with raw feelings, lies an immersive symphony waiting to be embraced. Constructed by an outstanding musician of language, this fascinating masterpiece conducts readers on a psychological trip, skillfully unraveling the hidden tunes and profound influence resonating within each carefully constructed phrase. Within the depths of this moving examination, we will examine the book's key harmonies, analyze its enthralling publishing fashion, and submit ourselves to the profound resonance that echoes in the depths of readers' souls.

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