

Practice Problems Thermochemistry chapter 5

1. The specific heat of liquid bromine is $0.750 \text{ J/g}\cdot\text{K}$. How much heat (J) is required to raise the temperature of 10.0 mL of bromine from 25.00°C to 35.0°C ? The density of liquid bromine is 2.1 g/mL .

Note: the specific heat of 0.750 is bogus. I pulled it out of the air. However, the problem still solves the same way.

2. The value of ΔH° for the reaction below is -6535 kJ . _____ kJ of heat are released in the combustion of 16.0 g of C_6H_6 (l)?



3. Carbon monoxide and oxygen gas react to form carbon dioxide. How much heat is released when 89.5 grams of O_2 (g) reacts with excess CO? $\Delta H^\circ = -482 \text{ kJ}$.



4. The value of ΔH° for the reaction below is -336 kJ . Calculate the heat (kJ) released to the surroundings when 23.0 g of HCl is formed.



5. How much heat is released when 29.5 grams of Cl_2 (g) reacts with excess hydrogen?



Thermochemistry Equation Practice Problems

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