

/5 K

/5 A

/6 T

/4 C

Attendance # _____

Part A: Multiple Choice

Circle the best answer

1. Two reactants are mixed in a bomb calorimeter and a reaction occurs. The temperature of the calorimeter drops. Which of the following statements regarding the reaction is true?

- a) The system is releasing heat.
b) The system is absorbing heat.
 c) The heat of reaction is equal to $q\Delta T$.
 d) The reaction is exothermic.
 e) A and D are correct.

2. Which of the following are exothermic?

- I. boiling water II. freezing water
 III. condensing steam IV. melting ice

- a) I and II b) I and IV **c) II and III**
 d) II and IV e) II, III, and IV

3. Which of the following will show a negative entropy change in the forward direction?

- I. $\text{SO}_{2(g)} + \text{CaO}_{(s)} \rightarrow \text{CaSO}_{3(s)}$
 II. $\text{Ni}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{H}_{2(g)} + \text{NiCl}_{2(aq)}$
 III. hydrogen and oxygen combine to form water molecules
 IV. air is pumped into a tire
 V. sugar dissolves in water

- a) I, III and IV** b) II and V c) II and IV
 d) III and V e) V only

Part B: Application

Complete the questions in the spaces provided. Be sure to record your answers with proper significant digits and units: (5 marks)

1. The reduction of iron (II) oxide (FeO) to metallic iron (Fe) can be performed using methane (CH_4) according to the following reaction:



Given the following heats of formation, determine the heat required to produce exactly 250.g of iron. (5 marks)

$$\Delta H^\circ_f \text{FeO}_{(s)} = -238 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CH}_{4(g)} = -73.61 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CO}_{2(g)} = -393.5 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{H}_2\text{O}_{(g)} = -241.8 \text{ kJ/mol}$$

$$\Delta H = [4 \times 0 \text{ kJ/mol} + 1 \times -393.5 \text{ kJ/mol} + 2 \times -241.8 \text{ kJ/mol}] - [1 \times -73.61 \text{ kJ/mol} + 4 \times -238 \text{ kJ/mol}]$$

$$\Delta H = [-877.1 \text{ kJ/mol}] - [-1025.61 \text{ kJ/mol}]$$

$$\Delta H = 149 \text{ kJ/mol CH}_4 \rightarrow 37.25 \text{ kJ/mol Fe}$$

$$n = m/M$$

$$= 250 \text{ g} / 55.845 \text{ g/mol}$$

$$= 4.477 \text{ mol}$$

$$37.25 \text{ kJ/mol} \times 4.477 \text{ mol} = 167 \text{ kJ}$$

4. Which of the following is the correct thermochemical equation for the formation of $\text{C}_6\text{H}_5\text{OH}_{(l)}$, which absorbs 235.2 kJ/mol when formed?

- a) $2\text{C}_{(s)} + 5/2 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)} + 235.2 \text{ kJ}$
 b) $2\text{C}_{(s)} + 3 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)} + 235.2 \text{ kJ}$
 c) $2\text{C}_{(s)} + 5/2 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} + 235.2 \text{ kJ} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)}$
d) $2\text{C}_{(s)} + 3 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} + 235.2 \text{ kJ} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)}$
 e) $\text{C}_6\text{H}_5\text{OH}_{(l)} + 3\text{O}_{2(g)} \rightarrow 2\text{CO}_{2(g)} + 3\text{H}_2\text{O}_{(l)} + 235.2 \text{ kJ}$

5. When solid sodium hydroxide (NaOH) is added to water, the beaker feels warm to the touch. Which statement(s) best describe(s) this observation?

- I. $\text{NaOH}_{(s)} \rightarrow \text{NaOH}_{(aq)} + 43.6 \text{ kJ}$**
 II. $\text{NaOH}_{(s)} \rightarrow \text{NaOH}_{(aq)} \Delta H = +43.6 \text{ kJ}$
 III. The reaction is endothermic
 IV. Heat is absorbed by the system, so it feels warmer
 V. Heat is absorbed by the reactants in order to decrease the entropy of the system.

- a) I only** b) II & III c) II, III and IV
 d) II and V e) V only

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**Andrew Hunt, Graham Curtis, Graham
Hill**



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Table of Contents Thermochem Exam With Answers

1. Understanding the eBook Thermochem Exam With Answers
 - The Rise of Digital Reading Thermochem Exam With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochem Exam With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochem Exam With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochem Exam With Answers
 - Personalized Recommendations
 - Thermochem Exam With Answers User Reviews and Ratings
 - Thermochem Exam With Answers and Bestseller Lists
5. Accessing Thermochem Exam With Answers Free and Paid eBooks
 - Thermochem Exam With Answers Public Domain eBooks
 - Thermochem Exam With Answers eBook Subscription Services
 - Thermochem Exam With Answers Budget-Friendly Options
6. Navigating Thermochem Exam With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochem Exam With Answers Compatibility with Devices
 - Thermochem Exam With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochem Exam With Answers
 - Highlighting and Note-Taking Thermochem Exam With Answers
 - Interactive Elements Thermochem Exam With Answers
8. Staying Engaged with Thermochem Exam With Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Thermochem Exam With Answers
- 9. Balancing eBooks and Physical Books Thermochem Exam With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochem Exam With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochem Exam With Answers
 - Setting Reading Goals Thermochem Exam With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochem Exam With Answers
 - Fact-Checking eBook Content of Thermochem Exam With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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