Types of Chemical Reaction and Predicting Products Worksheet

Balance the following. Indicate which type of chemical reaction (synthesis, decomposition, singledisplacement, double-displacement, acid-base, combustion, precipitation, or gas evolution) is being represented in 1 to 13.

- Na₃PO₄ (aq)+ KOH(aq) → NaOH(aq) + K₃PO₄(s) Reaction Type
- MgCl_{2 (S)} + ___Li₂CO_{3(N)} → ___MgCO_{3(N)} + ___LiCl(aq) Reaction Type
- C₆H₁₂₍₁₎ + O_{2(q)} → CO_{2(q)} + H₂O(I) Reaction Type
- Pb(s) + FeSO_{4(sq)} → PbSO_{4(s)} + Fe(s) Reaction Type _____
- CaCO_{3(s)} → CaO(s) + CO₂ (g) Reaction Type
- P_{4(s)} + O_{2(g)} → P₂O_{3(s)} Reaction Type _____
- RbNO_{3(aq)} + BeF_{2(aq)} → Be(NO₃)_{2(aq)} + RbF(aq) Reaction Type _____
- AgNO_{3(aq)} + Cu(s) → Cu(NO₃)_{2(aq)} + Ag(s) Reaction Type _____
- 9. $C_3H_6O(I) + O_{2(g)} \rightarrow CO_{2(g)} + H_2O(I)$
- C₅H₅₍₀ + Fe(S) → Fe(C₅H₅)₂ (S) Reaction Type _____
- SeCl₆ (s)+ O_{2(g)} → SeO_{2(s)} + Cl_{2(g)}
- Reaction Type _____
- MgI_{2(aq)} + Mn(SO₃)_{2(s)} → MgSO_{3(s)} + MnI₄(aq) Reaction Type
- NO_{2(g)} → O_{2(g)} + N_{2(g)}

Reaction Type

Reaction Type _____

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