

Unit 1 Chemistry Experiment: Writing balanced equations

Aim: To develop chemical equation writing skills.

Background

Writing balanced equations to represent or predict chemical reactions is a fundamental skill for chemists. A systematic approach helps with this process. In particular,

- always refer to **electrovalency** charts before writing formulas.

TABLE 5.5.1 ELECTROVALENCIES OF SOME COMMON CATIONS

+1	+2	+3	+4
Lithium Li^+ Sodium Na^+ Potassium K^+ Hydrogen H^+ Copper(I) Cu^+ Silver Ag^+ Ammonium NH_4^+ Hydronium H_3O^+	Magnesium Mg^{2+} Calcium Ca^{2+} Strontium Sr^{2+} Barium Ba^{2+} Iron(II) Fe^{2+} Copper(II) Cu^{2+} Nickel Ni^{2+} Zinc Zn^{2+} Tin(II) Sn^{2+} Lead(II) Pb^{2+} Manganese(II) Mn^{2+}	Aluminium Al^{3+} Chromium(III) Cr^{3+} Iron(III) Fe^{3+}	Tin(IV) Sn^{4+} Lead(IV) Pb^{4+}

TABLE 5.5.2 ELECTROVALENCIES OF SOME COMMON ANIONS

-1	-2	-3
Fluoride F^- Chloride Cl^- Bromide Br^- Iodide I^- Hydride H^- Hydroxide OH^- Nitrate NO_3^- Nitrite NO_2^- Permanganate MnO_4^- Chlorate ClO_3^- Cyanide CN^- Hydrogen sulfate HSO_4^- Hydrogen sulfite HSO_3^- Hydrogen sulfide HS^- Hydrogen carbonate HCO_3^-	Oxide O^{2-} Peroxide O_2^{2-} Sulfate SO_4^{2-} Sulfite SO_3^{2-} Sulfide S^{2-} Carbonate CO_3^{2-} Chromate CrO_4^{2-} Dichromate $\text{Cr}_2\text{O}_7^{2-}$ Thiosulfate $\text{S}_2\text{O}_3^{2-}$ Monohydrogen orthophosphate HPO_4^{2-}	Nitride N^{3-} Orthophosphate PO_4^{3-}

- reaction products can often be predicted by swapping the ions.

Examples



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