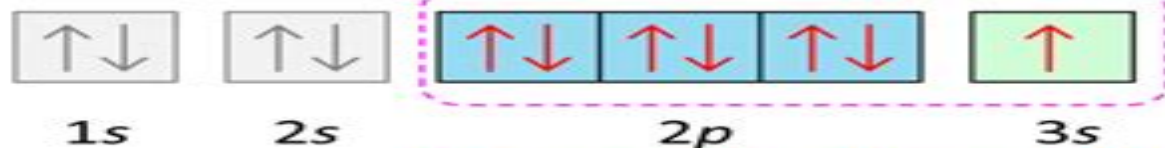


Electron Configuration of Main Group Cations

Electron configuration

Na $1s^2 2s^2 2p^6 3s^1$

Orbital diagram

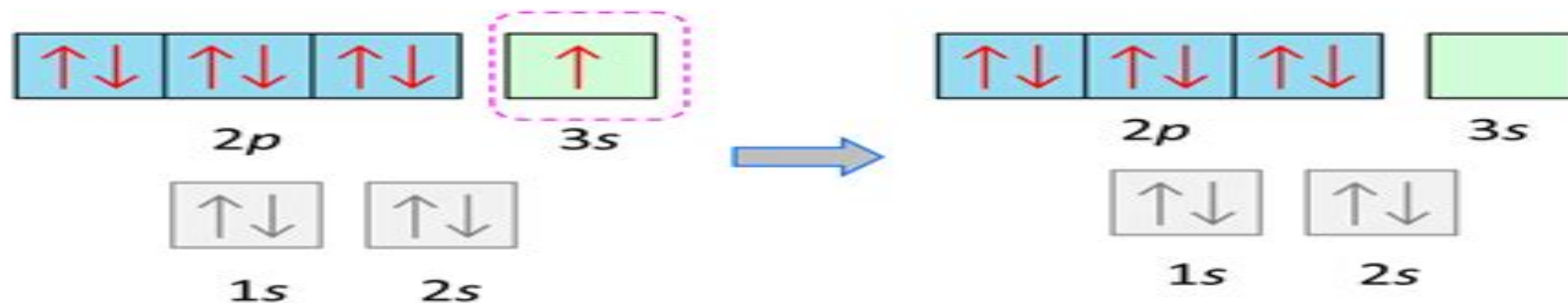


Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital – $3s^1$



Na $1s^2 2s^2 2p^6 3s^1$

Na⁺ $1s^2 2s^2 2p^6$

Same as
↓

Ne $1s^2 2s^2 2p^6$



ChemistrySteps.com

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

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Ying Liu



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Table of Contents Small Scale Lab Electoron Configurations Of Ions

1. Understanding the eBook Small Scale Lab Electoron Configurations Of Ions
 - The Rise of Digital Reading Small Scale Lab Electoron Configurations Of Ions
 - Advantages of eBooks Over Traditional Books
2. Identifying Small Scale Lab Electoron Configurations Of Ions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Small Scale Lab Electoron Configurations Of Ions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Small Scale Lab Electoron Configurations Of Ions
 - Personalized Recommendations
 - Small Scale Lab Electoron Configurations Of Ions User Reviews and Ratings

- Small Scale Lab Electoron Configurations Of Ions and Bestseller Lists
- 5. Accessing Small Scale Lab Electoron Configurations Of Ions Free and Paid eBooks
 - Small Scale Lab Electoron Configurations Of Ions Public Domain eBooks
 - Small Scale Lab Electoron Configurations Of Ions eBook Subscription Services
 - Small Scale Lab Electoron Configurations Of Ions Budget-Friendly Options
- 6. Navigating Small Scale Lab Electoron Configurations Of Ions eBook Formats
 - ePub, PDF, MOBI, and More
 - Small Scale Lab Electoron Configurations Of Ions Compatibility with Devices
 - Small Scale Lab Electoron Configurations Of Ions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Small Scale Lab Electoron Configurations Of Ions
 - Highlighting and Note-Taking Small Scale Lab Electoron Configurations Of Ions
 - Interactive Elements Small Scale Lab Electoron Configurations Of Ions
- 8. Staying Engaged with Small Scale Lab Electoron Configurations Of Ions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Small Scale Lab Electoron Configurations Of Ions
- 9. Balancing eBooks and Physical Books Small Scale Lab Electoron Configurations Of Ions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Small Scale Lab Electoron Configurations Of Ions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Small Scale Lab Electoron Configurations Of Ions
 - Setting Reading Goals Small Scale Lab Electoron Configurations Of Ions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Small Scale Lab Electoron Configurations Of Ions
 - Fact-Checking eBook Content of Small Scale Lab Electoron Configurations Of Ions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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