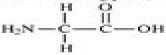
## Ch 14 Practice MC problems

- Which of the following molecules contains a nitrogen atom that is sp<sup>2</sup> hybridized?
  - A) NH<sub>2</sub>
  - B) NO<sub>3</sub>:
  - C) N<sub>2</sub>
  - D) HCN
  - E) C<sub>0</sub>N<sub>0</sub>
- 2. Consider the structure of glycine, the simplest amino acid:



What is the total number of  $\pi$  bonds in the molecule?

- A0 = 0
- B) 172.
- C) 1
- D) 2
- E) More information is needed.
- What is the hybridization of I in the molecule ICh?
  - A) sc
    - B) sp<sup>2</sup>
  - C) sp<sup>3</sup>
  - D) dsp<sup>3</sup>
  - E) d'sp<sup>3</sup>
- 4. What is the hybridization of S in the molecule H<sub>2</sub>S?
  - A3 39
  - B) sp<sup>2</sup>
  - C) sp<sup>8</sup>
  - D) dsp<sup>3</sup>
  - E) d'sp'
- What is the hybridization of C in the ion CN?
  - A) sp
  - B) sp<sup>2</sup>
  - C) sp<sup>N</sup>
  - D) dsp<sup>b</sup>
  - E) d'sp'
- Atoms that are sp<sup>1</sup> hybridized form \_\_\_\_ pi bond(s).
  - A(3) = 0
  - B) 1
  - C) 2
  - D) 3
  - E) 4
- What is the hybridization of the central atom in SF<sub>6</sub>?
  - A) sp
  - B) sp<sup>2</sup>
  - C) sp<sup>3</sup>
  - D) dsp<sup>3</sup>
  - E)  $d^2sp^3$

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**Angelica M. Stacy** 

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