# Chapter 12 Review 2

#### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

	mutual electrical attraction be		uclei and valence electrons of different atoms that
40.	dipole.	e.	chemical bond.
b.	Lewis structure.	d.	London force.
2Th	e electrons involved in the for	rmation of a	chemical bond are called
. 0.	dipoles.	C.	Lewis electrons.
ъ.	s electrons.	d.	valence electrons.
	chemical bond resulting from lled a(n)	the electros	tatic attraction between positive and negative ions is
20.		42.5	charged bond.
ь.	ionic bond.		dipole bond.
400	two covalently bonded atoms	are identica	I, the bond is
. 0.	nonpolar covalent.		nonionic.
b.	polar covalent.	d.	coordinate covalent.
510	the atoms that share electrons	have an une	equal attraction for the electrons, the bond is called
. a.	nonpolar.		ionic.
ъ.	polar.	d.	dipolar.
6A.	covalent bond results when	are shar	ed.
. 44.	ions	e.	electrons
b.	Lewis structures	d.	dipoles
7Ne	enpolar covalent bonds are not	common b	ecause
. 40.	one atom usually attracts electrons more strongly than the other.		
Ъ.	ions always form when atoms join.		
e.	the electrons usually remain equally distant from both atoms.		
d.	dipoles are rare in nature.		
8Th	e greater the electronegativity	difference	between two bonded atoms, the greater the percentage
0.0	ionic character.	C.	metallic character.
Ь.	covalent character.	d.	electron sharing.
9Th	e pair of elements that forms	a bond with	the least ionic character is
. 40.	Na and Cl.	e.	O and Cl.
ь.	H and Cl.	d.	Br and Cl.
0. In	which of these compounds is	the bond be	etween the atoms NOT a nonpolar covalent bond?
40	Cl <sub>2</sub>	e.	HCI
b.	H <sub>2</sub>	di.	O <sub>2</sub>

# **Standardized Test Practice Chapter 3 Chemistry**

Carl A. Burtis, Edward R. Ashwood, David E. Bruns

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