

ANSWER KEY-Chem 200-Final Exam sample

7	A buffer solution contains 0.45 M sodium acetate (CH_3COONa) and 0.36 M acetic acid (CH_3COOH), $\text{pK}_a = 4.75$. What is the pH of this solution? A. 5.21 B. 4.65 C. 4.85 D. 3.60
8	Which of the following acids is the STRONGEST? A. HF ($\text{pK}_a = 3.14$) B. HNO_2 ($\text{pK}_a = 3.35$) C. HCN ($\text{pK}_a = 9.4$) D. HCOOH ($\text{pK}_a = 3.74$)
10	The following acids are listed with their pK_a values. Arrange the conjugate bases in order of INCREASING basic strength: <u>Conjugate acids:</u> NH_4^+ $(\text{CH}_3)_3\text{NH}^+$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_3^+$ $(\text{CH}_3\text{CH}_2)_3\text{NH}^+$ <u>pK_a:</u> 9.2 9.8 10.5 10.7 A. $\text{NH}_3 < (\text{CH}_3)_3\text{N} < \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 < (\text{CH}_3\text{CH}_2)_3\text{N}$ B. $\text{NH}_3 < \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 < (\text{CH}_3)_3\text{N} < (\text{CH}_3\text{CH}_2)_3\text{N}$ C. $(\text{CH}_3\text{CH}_2)_3\text{N} < \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 < (\text{CH}_3)_3\text{N} < \text{NH}_3$ D. $(\text{CH}_3\text{CH}_2)_3\text{N} < (\text{CH}_3)_3\text{N} < \text{NH}_3 < \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2$
11	In order to detect the equivalence point in a titration, the chosen colored indicator should have a pK_a equal to A. pH at equivalence B. pH at equivalence ± 1 C. pH at equivalence + 1 D. pH at equivalence - 1
12	Identify the reactant that is a BRÖNSTED-LOWRY ACID in the following reaction: $\text{KCN} + \text{HI} \leftrightarrow \text{HCN} + \text{KI}$ A. KCN B. HI C. HCN D. KI
13	Which of the following statements is incorrect at 25°C? A. $\text{pH} = -\log[\text{H}^+]$ B. $\text{pOH} = -\log[\text{OH}^-]$ C. $\text{pH} + \text{pOH} = 16$ D. $\text{pH} + \text{pOH} = \text{pK}_w$
14	What is TRUE about PURE water? A. $\text{pH} = 7$ at all temperatures B. $[\text{H}^+] = [\text{OH}^-]$ at all temperatures C. $K_w = 10^{-14}$ all temperatures D. All of the above

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Michael Rajnik

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