

Answer to Mole Problems**Mole Worksheet KEY**

Name _____

Period _____

Molar Relationships

Directions: Answer the following questions **on a separate sheet of paper**. Set-up all problems using the factor-label method of dimensional analysis and show all your work and units.

1. What is the mass of 7.50 moles of sulfur dioxide (SO_2)? **480g**

$$\frac{7.5\text{mol}}{1} \times \frac{64\text{g}}{1\text{mol}} = 480\text{g}$$

2. How many moles are there in 250.0 grams of sodium phosphate (Na_3PO_4)? **1.52mol**

$$\frac{250\text{g}}{1} \times \frac{1\text{mol}}{164\text{g}} = 1.52\text{mol}$$

3. How many grams of potassium sulfate (K_2SO_4) are there in 25.3 moles? **4402g**

$$\frac{25.3\text{mol}}{1} \times \frac{176\text{g}}{1\text{mol}} = 4402\text{g}$$

4. What is the volume of 0.38 moles of any gas at STP? **8.5L**

$$\frac{0.38\text{mol}}{1} \times \frac{22.4\text{L}}{1\text{mol}} = 8.5\text{L}$$

7. What is the mass of 51 liters of oxygen gas? **73g**

$$\frac{51\text{L}}{1} \times \frac{1\text{mol}}{22.4\text{L}} \times \frac{32\text{g}}{1\text{mol}} = 73\text{g}$$

1. Calculate the number of moles in 32.2-L of NH_3

$$\frac{32.3\text{L}}{1} \times \frac{1\text{mol}}{22.4\text{L}} = 1.44\text{mol}$$

2. Calculate the number of grams in 3.25-mol of AgNO_3

Using Moles With Formulas Packet Answers

Roger Frost



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