

		electrons can move from valence band to conduction band. Or, the diagram.				
		(a) Metal (b) Insulator				
19.		The preparation of Potassium permanganate from Pyrolusite (MnO_2) involves two steps. 1. MnO_2 is fused with KOH to form potassium manganate (K_2MnO_4). $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 \rightarrow 2\text{K}_2\text{MnO}_4 + 2\text{H}_2\text{O}$ 2. K_2MnO_4 is electrolytically oxidised to potassium permanganate. $\text{MnO}_4^{2-} \xrightarrow[\text{in alkaline medium}]{\text{Electrolytic oxidation}} \text{MnO}_4^-$	1 1	2		
20.		1. When aniline is treated with nitrous acid (prepared by mixing NaNO_2 & HCl) at 273-278K, benzene diazonium chloride is formed. 2. Benzene diazonium chloride on warming with water to form phenol. Or the equation: Aniline Benzene diazonium chloride	1 1	2		
PART III A. Answer any 3 questions from 21 to 24. Each carries 3 scores						
21.		<table><tr><td><i>Schottky Defect</i> <i>Arising due to the missing of equal number of anions and cations from the lattice site</i> <i>Decreases the density of the solid</i> <i>It is shown by ionic crystals in which the anionic and cationic sizes are almost equal.</i></td><td><i>Frenkel Defect</i> <i>Arising due to the misplacing of a cation from the lattice site to the interstitial site.</i> <i>No change in the density of the solid.</i> <i>It is shown by ionic solids in which there is a large difference in the size of the ions.</i></td></tr></table>	<i>Schottky Defect</i> <i>Arising due to the missing of equal number of anions and cations from the lattice site</i> <i>Decreases the density of the solid</i> <i>It is shown by ionic crystals in which the anionic and cationic sizes are almost equal.</i>	<i>Frenkel Defect</i> <i>Arising due to the misplacing of a cation from the lattice site to the interstitial site.</i> <i>No change in the density of the solid.</i> <i>It is shown by ionic solids in which there is a large difference in the size of the ions.</i>	3	3
<i>Schottky Defect</i> <i>Arising due to the missing of equal number of anions and cations from the lattice site</i> <i>Decreases the density of the solid</i> <i>It is shown by ionic crystals in which the anionic and cationic sizes are almost equal.</i>	<i>Frenkel Defect</i> <i>Arising due to the misplacing of a cation from the lattice site to the interstitial site.</i> <i>No change in the density of the solid.</i> <i>It is shown by ionic solids in which there is a large difference in the size of the ions.</i>					
22.		For a first order reaction, $k = \frac{2.303}{t} \log \frac{[R]_0}{[R]}$	1	3		

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