- (5 points) Table 4.8 (<u>Snoeyink</u> and Jenkins) shows the concentration of carbonate species for solutions with distilled water and 10-5M solutions of H₂CO₃*; sodium bicarbonate, NaHCO₃; and sodium carbonate, Na₂CO₃, in a system open to the atmosphere and in a closed system.
 - a. Why does the concentration of H₂CO₃* remain constant in the open system (across the different solutions) and decreases in a closed system?
 - b. For both systems, the concentration of H* decreases when comparing a solution of distilled water and 10-5M solutions of H2CO3*; NaHCO3 Na2CO3 with lowest concentration observed for Na2CO3 The decrease in H* concentration is greater for a closed system than for an open system. Explain this difference.

TABLE 4.8 Comparison of Composition and pH of Open and Closed Carbonate Systems

Solution	Closed System (Fig. 4-17)						
	[H*]	[OH-]	[H _i CO;]	[HCO,"]	[CO,1-]	pН	C1,00
10-4 M H,CO;	10-4.7	10-4.9	10-41	10-tr	10-16.4	5.7	10-54
10" M NaHCO,	10-74	10-44	10-44	10-1	10-54	7.6	10-54
10-4 M Na,CO,	10-*	10-	10-7.4	10-5	10-4.5	9.0	10-44
Solution	Open System						
	[H*]	[OH-]	[H,CO;]	[HCO,-]	[CO,*-]	pН	C1.00,
10-4 M H,CO;	10-6.7	10-43	10-4	10-4.7	10-16.4	5.7	10-40
10" M NaHCO,	10-43	10 17	10-	10-4	10-44	6.3	10-4
10-4 M Na,00,	10-4.7	10-7.3	10-1	10-47	10-4-1	6.6	10-44

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Vernon Snoeynik

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