

17

WATER AND AQUEOUS SYSTEMS

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SECTION 17.1 LIQUID WATER AND ITS PROPERTIES (pages 475–478)

This section explains the hydrogen bonding that occurs in water molecules and how high surface tension and low vapor pressure can be explained in terms of hydrogen bonding.

► The Water Molecule (pages 475–476)

1. What unique molecular compound is the foundation of all life on Earth?

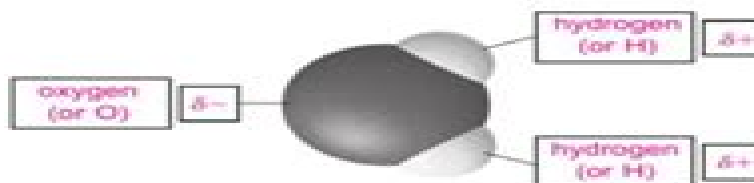
water

2. Approximately what fraction of Earth's surface is covered in water? 3/4

3. Circle the letter next to each sentence that is true concerning water molecules.

- a. Each O — H covalent bond in a water molecule is nonpolar.
☒ b. In a water molecule, the less electronegative hydrogen atoms acquire a slight positive charge and the oxygen atom acquires a slight negative charge.
☒ c. Because the water molecule has an H — O — H bond angle of 105° , the molecule as a whole is polar.

4. The diagram below depicts a water molecule. Complete the labels showing the locations of the hydrogen atoms, the oxygen atom, and the regions of positive and negative charge.



5. List five important properties of water that result from hydrogen bonding.

- a. high surface tension
 b. low vapor pressure
 c. high specific heat capacity
 d. high heat of vaporization
 e. high boiling point

Water And Aqueous Systems Guide Answers

M Tight



Water And Aqueous Systems Guide Answers:

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Table of Contents Water And Aqueous Systems Guide Answers

1. Understanding the eBook Water And Aqueous Systems Guide Answers
 - The Rise of Digital Reading Water And Aqueous Systems Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Water And Aqueous Systems Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Water And Aqueous Systems Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Water And Aqueous Systems Guide Answers
 - Personalized Recommendations
 - Water And Aqueous Systems Guide Answers User Reviews and Ratings
 - Water And Aqueous Systems Guide Answers and Bestseller Lists
5. Accessing Water And Aqueous Systems Guide Answers Free and Paid eBooks
 - Water And Aqueous Systems Guide Answers Public Domain eBooks
 - Water And Aqueous Systems Guide Answers eBook Subscription Services
 - Water And Aqueous Systems Guide Answers Budget-Friendly Options
6. Navigating Water And Aqueous Systems Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Water And Aqueous Systems Guide Answers Compatibility with Devices
 - Water And Aqueous Systems Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Water And Aqueous Systems Guide Answers
 - Highlighting and Note-Taking Water And Aqueous Systems Guide Answers
 - Interactive Elements Water And Aqueous Systems Guide Answers

8. Staying Engaged with Water And Aqueous Systems Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Water And Aqueous Systems Guide Answers
9. Balancing eBooks and Physical Books Water And Aqueous Systems Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Water And Aqueous Systems Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Water And Aqueous Systems Guide Answers
 - Setting Reading Goals Water And Aqueous Systems Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Water And Aqueous Systems Guide Answers
 - Fact-Checking eBook Content of Water And Aqueous Systems Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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