

SCH4U1 Unit 5 - Thermochemistry and Rates of Reaction
Summative Assessment

Name: Answer keys

K/U:

17

TH:

10

C:

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APP:

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1. Please answer the following questions in the space provided. [K5]

a) When solid KBr is dissolved in water, the solution gets colder. Is the chemical system exothermic or endothermic? [1]

endothermic

b) When a sample of potassium nitrate is dissolved in water, the temperature of the water changes from 23.5°C to 20.3°C. Is the chemical system exothermic or endothermic? [1]

endothermic

c) In the reaction $2A + B \rightarrow C + 3D$, reactant A is found to disappear at the rate of $6.2 \times 10^{-4} \text{ mol/L.s}$. What is the rate of disappearance of B? [1]

2 mol of A disappear at rate $6.2 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$

1 mol of B disappears at $x \frac{\text{mol}}{\text{L.s}}$
 $x = \frac{6.2 \times 10^{-4} \text{ L.s}}{2} = 3.1 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$

d) The following diagram shows a collision between two BrNO molecules. Use collision theory to explain the results of the collision. [2]



Reaction did take place. New compounds formed; 2 NO and 1 Br₂ are the products. It was an effective collision. Reactants collided with proper orientation and sufficient (optimal) kinetic energy.

2. Consider the following reaction:

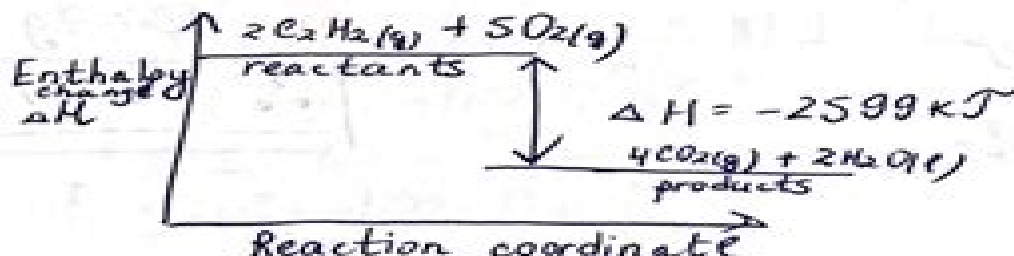


$$\Delta H = -2599 \text{ kJ}$$

a) Write a thermochemical equation with the enthalpy change, ΔH , 'inside' the equation. [C1]



b) Sketch an enthalpy diagram for this reaction. Provide appropriate labels for each axis. [C3]



Solution And Thermochemistry Test Answer Keys

Gary S. Thorpe



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Chemistry Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, 2004-07-01 Prentice Hall Chemistry meets the needs of students with a range of abilities diversities and learning styles by providing real world connections to chemical concepts and processes The first nine chapters introduce students to

the conceptual nature of chemistry before they encounter the more rigorous mathematical models and concepts in later chapters. The technology backbone of the program is the widely praised Interactive Textbook with ChemASAP which provides frequent opportunities to practice and reinforce key concepts with Animations Simulations Assessment and Problem solving tutorials

Sterling Test Prep MCAT General Chemistry Review: Complete Subject Review Sterling Test Prep, 2017-05-21

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Answer Key and Test Manual Chemistry for the Health Sciences George I. Sackheim, Dennis D. Lehman, Ronald M. Schultz, 1977

A Level Chemistry MCQs Arshad Iqbal, 2017-04-20

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