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CHAPTER 2

ATOMS, MOLECULES, AND IONS

Questions

18. Some elements exist as molecular substances. That is, hydrogen normally exists as H_2 molecules, not single hydrogen atoms. The same is true for N_2 , O_2 , F_2 , Cl_2 , etc.
19. A compound will always contain the same numbers (and types) of atoms. A given amount of hydrogen will react only with a specific amount of oxygen. Any excess oxygen will remain unreacted.
20. The halogens have a high affinity for electrons, and one important way they react is to form anions of the type X^- . The alkali metals tend to give up electrons easily and in most of their compounds exist as M^+ cations. *Note:* These two very reactive groups are only one electron away (in the periodic table) from the least reactive family of elements, the noble gases.
21. Law of conservation of mass: Mass is neither created nor destroyed. The total mass before a chemical reaction always equals the total mass after a chemical reaction.

Law of definite proportion: A given compound always contains exactly the same proportion of elements by mass. For example, water is always 1 g H for every 8 g oxygen.

Law of multiple proportions: When two elements form a series of compounds, the ratios of the mass of the second element that combine with 1 g of the first element always can be reduced to small whole numbers. For CO_2 and CO discussed in Section 2.2, the mass ratios of oxygen that react with 1 g carbon in each compound are in a 2 : 1 ratio.

22.
 - a. The smaller parts are electrons and the nucleus. The nucleus is broken down into protons and neutrons, which can be broken down into quarks. For our purpose, electrons, neutrons, and protons are the key smaller parts of an atom.
 - b. All atoms of hydrogen have 1 proton in the nucleus. Different isotopes of hydrogen have 0, 1, or 2 neutrons in the nucleus. Because we are talking about atoms, this implies a neutral charge, which dictates 1 electron present for all hydrogen atoms. If charged ions were included, then different ions/atoms of H could have different numbers of electrons.
 - c. Hydrogen atoms always have 1 proton in the nucleus, and helium atoms always have 2 protons in the nucleus. The number of neutrons can be the same for a hydrogen atom and a helium atom. Tritium (3H) and 4He both have 2 neutrons. Assuming neutral atoms, then the number of electrons will be 1 for hydrogen and 2 for helium.

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1. Understanding the eBook Solution Manual For Chemistry Zumdahl
 - The Rise of Digital Reading Solution Manual For Chemistry Zumdahl
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual For Chemistry Zumdahl
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual For Chemistry Zumdahl
 - User-Friendly Interface

4. Exploring eBook Recommendations from Solution Manual For Chemistry Zumdahl
 - Personalized Recommendations
 - Solution Manual For Chemistry Zumdahl User Reviews and Ratings
 - Solution Manual For Chemistry Zumdahl and Bestseller Lists
5. Accessing Solution Manual For Chemistry Zumdahl Free and Paid eBooks
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 - Solution Manual For Chemistry Zumdahl Enhanced eBook Features
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 - Highlighting and Note-Taking Solution Manual For Chemistry Zumdahl
 - Interactive Elements Solution Manual For Chemistry Zumdahl
8. Staying Engaged with Solution Manual For Chemistry Zumdahl
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual For Chemistry Zumdahl
9. Balancing eBooks and Physical Books Solution Manual For Chemistry Zumdahl
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual For Chemistry Zumdahl
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solution Manual For Chemistry Zumdahl
 - Setting Reading Goals Solution Manual For Chemistry Zumdahl
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Solution Manual For Chemistry Zumdahl
 - Fact-Checking eBook Content of Solution Manual For Chemistry Zumdahl
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
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