

SOLUTION & COLLIGATIVE PROPERTIES

OSMOTIC PRESSURE :

- (i) $\pi = \rho gh$ Where, ρ = density of soln., h = equilibrium height.
 (ii) **Vant – Hoff Formula** (For calculation of O.P.)
 $\pi = CST$

$$\pi = CRT = \frac{n}{V} RT \text{ (just like ideal gas equation)}$$

$$\begin{aligned} \therefore C &= \text{total conc. of all types of particles.} \\ &= C_1 + C_2 + C_3 + \dots \\ &= \frac{(n_1 + n_2 + n_3 + \dots)}{V} \end{aligned}$$

Note : If V_1 mL of C_1 conc. + V_2 mL of C_2 conc. are mixed.

$$\pi = \left(\frac{C_1 V_1 + C_2 V_2}{V_1 + V_2} \right) RT \quad ; \quad \pi = \left(\frac{\pi_1 V_1 + \pi_2 V_2}{RT} \right)$$

Type of solutions :

(a) **Isotonic solution** – Two solutions having same O.P.

$$\pi_1 = \pi_2 \text{ (at same temp.)}$$

(b) **Hyper tonic**– If $\pi_1 > \pi_2 \Rightarrow$ 1st solution is hypertonic solution w.r.t. 2nd solution.

(c) **Hypotonic** – If 2nd solution is hypotonic w.r.t. 1st solution.

Abnormal Colligative Properties : (In case of association or dissociation)

VANT HOFF CORRECTION FACTOR (i) :

$$\begin{aligned} i &= \frac{\text{exp/observed/actual/abnormal value of colligative property}}{\text{Theoretical value of colligative property}} \\ &= \frac{\text{exp./observed no. of particles / conc.}}{\text{Theoretical no. of particles}} = \frac{\text{observed molality}}{\text{Theoretical molality}} \\ &= \frac{\text{theoretical molar mass(formula mass)}}{\text{experimental/observed molar mass(apparent molar mass)}} \end{aligned}$$

- $i > 1 \Rightarrow$ dissociation.
 ○ $i < 1 \Rightarrow$ association.

○
$$i = \frac{\pi_{\text{exp.}}}{\pi_{\text{theor}}}$$

$$\begin{aligned} \therefore \pi &= iCRT \\ \pi &= (i_1 C_1 + i_2 C_2 + i_3 C_3, \dots) RT \end{aligned}$$

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CO Houle



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