CHAPTER 18 ACID-BASE EQUILIBRIA

FOLLOW-UP PROBLEMS

18.1A Plan: Examine the formulas and classify each as an acid or base. Strong acids are the hydrohalic acids HCl, HBr, and HI, and oxoacids in which the number of O atoms exceeds the number of ionizable protons by at least two. Other acids are weak acids. Strong bases are soluble oxides or hydroxides of the Group 1A(1) metals and Ca, Sr, and Ba in Group 2A(2). Other bases are weak bases.

a) Chloric acid, HCiO₃, is the stronger acid because acid strength increases as the number of O atoms in the acid increases.

b) Hydrochloric acid, HCl, is one of the strong hydrohalic acids whereas acetic acid, CH₃COOH, is a weak carboxylic acid.

c) Sodium hydroxide, NaOH, is a strong base because Na is a Group 1A(1) metal. Methylamine, CH₃NH₂, is an organic amine and, therefore, a weak base.

18.1B Plan: Examine the formulas and classify each as a strong acid, weak acid, strong base, or weak base. Strong acids are the hydrohalic acids HCL HBr, and HL and oxoacids in which the number of O atoms exceeds the number of ionizable protons by at least two. Other acids are weak acids. Strong bases are soluble oxides or hydroxides of the Group 1A(1) metals and Ca, Sr, and Ba in Group 2A(2). Other bases are weak bases.
Solution:

a) (CH₃)₃N is a weak base. It contains a nitrogen atom with a lone pair of electrons, which classifies it as a base; however, it is not one of the strong bases.

b) Hydroiodic acid. HI, is a strong acid (one of the strong acids listed above).

c) HBrO is a weak acid. It has an ionizable hydrogen, which makes it an acid. Specifically, it is an oxoacid, in which a polyatomic ion is the anion. In the case of this oxoacid, there is only one O atom for each ionizable hydrogen, so this is a weak acid. The reaction for the dissociation of this weak acid is:

 $HBrO(\omega_q) + H_2O(t) \Rightarrow BrO'(\omega_q) + H_3O'(\omega_q)$. The corresponding equilibrium expression is:

$$K_{\pm} = \frac{\left[B_{\ell}O^{\ell}\right][H_{3}O^{\ell}]}{[HB_{\ell}O]}$$

d) Ca(OH)₂ is a strong base (one of the strong bases listed above).

18.2A Plan: The product of [H₃O*] and [OH*] remains constant at 25°C because the value of K_n is constant at a given temperature. Use K_n = [H₃O*][OH*] = 1.0 x 10⁻¹⁴ to solve for [H₃O*].
Solution:

Calculating [H₂O"]:

$$[\mathrm{H}_{3}\mathrm{O}^{*}] = \frac{K_{w}}{[\mathrm{OH}^{*}]} = \frac{1.0 \times 10^{-14}}{6.7 \times 10^{-2}} = 1.4925 \times 10^{-13} = 1.5 \times 10^{-13} \, M$$

Since $[OH^*] > [H_1O^*]$, the solution is basic.

18.2B Plan: The product of [H₂O*] and [OH*] remains constant at 25°C because the value of K_w is constant at a given temperature. Use K_w = [H₂O*][OH*] = 1.0 x 10⁻¹⁴ to solve for [H₂O*].
Solution:

Solution: Calculating [OH]:

$$[OH^-] = \frac{K_\infty}{[H_1O^+]} = \frac{1.0 \times 10^{-0.4}}{1.8 \times 10^{-0.0}} = 5.55555 \times 10^{-3} = 5.6 \times 10^{-8} M$$

Since $[OH^*] > [H_3O^*]$, the solution is basic.

18.3A Plan: NaOH is a strong base that dissociates completely in water. Subtract pH from 14.00 to find the pOH, and calculate inverse logs of pH and pOH to find [H₁O*] and [OH], respectively.

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kinetic energy and speed histograms for light and heavy particles explore diffusion and determine how concentration temperature mass and radius affect the rate of diffusion

gas properties gas heat thermodynamics phet interactive simulations - Aug 12 2023

web translate this sim related simulations balloons buoyancy states of matter states of matter basics software requirements credits pump gas molecules to a box and see what happens as you change the volume add or remove heat change gravity and more

gas properties modular homework activity phet contribution - Jun 10 2023

web may 6 2013 this activity has 5 modules explore the simulation kinetic energy and speed kinetic molecular theory of gases relationships between gas variables pressure and mixtures of gases after exploring the simulation students can do the other modules in any order subject chemistry

gases intro ideal gas law phet interactive simulations - Apr 27 2022

web measure the temperature and pressure and discover how the properties of the gas vary in relation to each other pump gas molecules to a box and see what happens as you change the volume add or remove heat and more

gas properties simulation phet contribution - Feb 23 2022

web dec 3 2019 gas properties simulation description this is a guided inquiry lab in which students explore the relationships between pressure temperature volume and rates of diffusion students explore the relationships between volume and pressure temperature and kinetic energy amount of matter and pressure and temperature and pressure

virtual lab activity on gas properties phet contribution - Mar 27 2022

web mar 4 2022 the virtual lab activity on gas properties aims to describe the relationships of the different gas properties such as pressure volume temperature and amount of a substance to each other subject phet simulation gas properties aapt - Apr 08 2023

web aug 18 2016 this simulation provides an interactive simulation of an ideal gas users can change the number of gas molecules in a chamber change the volume add remove heat and change gravity users can explore the relationship between changes in kinetic energy of the gas molecules and heat