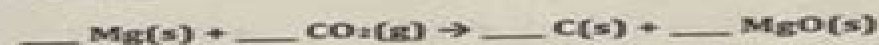


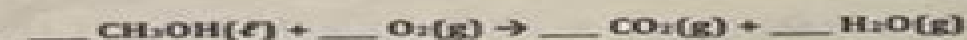
Limiting Reactant Problems



1. If 0.15 mol Al is reacted with 0.20 mol F_2 which is the limiting reactant and which is in excess?
3. How many mol AlF_3 will be produced?
4. What mass of AlF_3 will be obtained?
5. When I do the reaction I obtain 10.0 g AlF_3 as product, what is my % yield?



1. Balance the equation.
2. If 1.2 g Mg is reacted with 1.76 g CO_2 which is the limiting reactant and which is in excess?
3. What mass of MgO will be produced?
4. When I do the reaction I obtain 1.70 g MgO as product, what is my % yield?



1. Balance the equation.
2. If 6.40 g CH_3OH is reacted with 16.0 g O_2 which is the limiting reactant and which is in excess?
3. What mass of $\text{H}_2\text{O}(g)$ will be produced?



1. Balance the equation.
2. If 4.05 g Al is reacted with 14.2 g Cl_2 which is the limiting reactant and which is in excess?
3. What mass of AlCl_3 will be produced?
4. If I performed the reaction and obtained 10.0 g AlCl_3 as product, what is my % yield?

Ws 12 3 Limiting Reactant Problems

SA Dillow



Ws 12 3 Limiting Reactant Problems:

Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Ws 12 3 Limiting Reactant Problems**

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polaritäten des lebens sie berühren unser innerstes wesen dessen tiefe in ritualen erfahrbar ist