

## Chemical Formula Writing Worksheet Solutions

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write "ZnCl<sub>2</sub>", as shown.

|                 | <i>zinc</i>                                                    | <i>iron (II)</i>                                               | <i>iron (III)</i>                                              | <i>gallium</i>                                                 | <i>silver</i>                                     | <i>lead (IV)</i>                                               |
|-----------------|----------------------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------|---------------------------------------------------|----------------------------------------------------------------|
| <i>chloride</i> | <b>ZnCl<sub>2</sub></b>                                        | <b>FeCl<sub>2</sub></b>                                        | <b>FeCl<sub>3</sub></b>                                        | <b>GaCl<sub>3</sub></b>                                        | <b>AgCl</b>                                       | <b>PbCl<sub>4</sub></b>                                        |
| <i>acetate</i>  | <b>Zn(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> | <b>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> | <b>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></b> | <b>Ga(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></b> | <b>Ag C<sub>2</sub>H<sub>3</sub>O<sub>2</sub></b> | <b>Pb(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>4</sub></b> |
| <i>nitrate</i>  | <b>Zn(NO<sub>3</sub>)<sub>2</sub></b>                          | <b>Fe(NO<sub>3</sub>)<sub>2</sub></b>                          | <b>Fe(NO<sub>3</sub>)<sub>3</sub></b>                          | <b>Ga(NO<sub>3</sub>)<sub>3</sub></b>                          | <b>AgNO<sub>3</sub></b>                           | <b>Pb(NO<sub>3</sub>)<sub>4</sub></b>                          |
| <i>oxide</i>    | <b>ZnO</b>                                                     | <b>FeO</b>                                                     | <b>Fe<sub>2</sub>O<sub>3</sub></b>                             | <b>Ga<sub>2</sub>O<sub>3</sub></b>                             | <b>Ag<sub>2</sub>O</b>                            | <b>PbO<sub>2</sub></b>                                         |
| <i>nitride</i>  | <b>Zn<sub>3</sub>N<sub>2</sub></b>                             | <b>Fe<sub>3</sub>N<sub>2</sub></b>                             | <b>FeN</b>                                                     | <b>GaN</b>                                                     | <b>Ag<sub>3</sub>N</b>                            | <b>Pb<sub>3</sub>N<sub>4</sub></b>                             |
| <i>sulfate</i>  | <b>ZnSO<sub>4</sub></b>                                        | <b>FeSO<sub>4</sub></b>                                        | <b>Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>              | <b>Ga<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>              | <b>Ag<sub>2</sub>SO<sub>4</sub></b>               | <b>Pb(SO<sub>4</sub>)<sub>2</sub></b>                          |

Write the formulas for the following compounds:

- 1) copper (II) chloride **CuCl<sub>2</sub>**
- 2) lithium acetate **LiC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>**
- 3) vanadium (III) selenide **V<sub>2</sub>Se<sub>3</sub>**
- 4) manganese (IV) nitride **Mn<sub>3</sub>N<sub>4</sub>**
- 5) beryllium oxide **BeO**
- 6) sodium sulfate **Na<sub>2</sub>SO<sub>4</sub>**
- 7) aluminum arsenide **AlAs**
- 8) potassium permanganate **KMnO<sub>4</sub>**
- 9) chromium (VI) cyanide **Cr(CN)<sub>6</sub>**
- 10) tin (II) sulfite **SnSO<sub>3</sub>**
- 11) vanadium (V) fluoride **VF<sub>5</sub>**
- 12) ammonium nitrate **NH<sub>4</sub>NO<sub>3</sub>**

# Writing Chemical Formulas Worksheet Key Practice Problems Level 2

**RM Cervero**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

## Writing Chemical Formulas Worksheet Key Practice Problems Level 2:

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## **Writing Chemical Formulas Worksheet Key Practice Problems Level 2** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the energy of words has be more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Writing Chemical Formulas Worksheet Key Practice Problems Level 2**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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### **Table of Contents Writing Chemical Formulas Worksheet Key Practice Problems Level 2**

1. Understanding the eBook Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - The Rise of Digital Reading Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Personalized Recommendations
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 User Reviews and Ratings
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 and Bestseller Lists

5. Accessing Writing Chemical Formulas Worksheet Key Practice Problems Level 2 Free and Paid eBooks
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 Public Domain eBooks
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 eBook Subscription Services
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 Budget-Friendly Options
6. Navigating Writing Chemical Formulas Worksheet Key Practice Problems Level 2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 Compatibility with Devices
  - Writing Chemical Formulas Worksheet Key Practice Problems Level 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Highlighting and Note-Taking Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Interactive Elements Writing Chemical Formulas Worksheet Key Practice Problems Level 2
8. Staying Engaged with Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Writing Chemical Formulas Worksheet Key Practice Problems Level 2
9. Balancing eBooks and Physical Books Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Writing Chemical Formulas Worksheet Key Practice Problems Level 2
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Setting Reading Goals Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Fact-Checking eBook Content of Writing Chemical Formulas Worksheet Key Practice Problems Level 2
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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