| NAME_ | Period Date |
|---------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| | Predicting Reaction types S, D, SD, DB, C 62 NR |
| equatio | the type of reaction (or NR for no reaction,) complete the word equation, and write a balanced chemical on for each of the following. Make sure to include phase symbols. → Wildel State of model of model of the following and hydrochloric acid → |
| DD | $N(s) + N(q) \rightarrow N(c)_{s}(q) + H_{2}(g)$ 2. Calcium hydroxide and nitric acid \Rightarrow |
| NR | Ca(0H)= (aq) + 2HNO ₅ (q) -> Ca(NO ₃) ₃ +2H ₂ O 3. Aluminum and Magnesium → |
| S | Al (s) + Mg(s) X> contract 2 materials goth tend to 4. Mercury(II) and oxygen > Form contract |
| DD. | $2 + 9 (t) + 0_2 (9) \longrightarrow 2 + 9 0 (s)$ 5. Zinc chloride and hydrogen sulfide \Rightarrow |
| 000 | $ZnCl_2(aq) + H_2S(q) \longrightarrow ZnS(s) + 2HCL(aq)$ 6. Silver chloride and sodium nitrate \rightarrow |
| Pyroli | 7. Sodium chlorate is heated > Na. NO3 (ag) X> both products are agricult most |
| DD | 2 NaClO 2(s) A > 2 NaCl/m)+3O2 (a) bc a prompitate, N2O, 8. Barium nitrate and sodium chromate → |
| DD | Ba(NO3) 2 (ag) + Nacr O4 (ag) -> Ba CrO4(s) +2NaNO3 |
| DD | NaBr (q) + AgNO $_3$ (aq) \rightarrow AgBr (s) + Na No $_3$ (aq) (eq) |
| D | $Ca_3(PO_4)_2(s) + Al_2(sO_4)_3(qq) \longrightarrow 2AlPO_4(s) + 3CoSO_4(s)$ 11. Zinc carbonate is heated \Rightarrow |
| | $Z_0 C_0 C_3 (s) \xrightarrow{A} Z_0 C_3 + C_0 C_3 + C_0 C_3 $ |
| | $2K(s) + F_2(g) \longrightarrow 2KF(s)_{map ?}$ 13. Hydrogen and oxygen \Rightarrow |
| | $2H_2(g) + O_2(g) \longrightarrow 2H_2O(g)$ 14. (lead(II) and potassium chlorate \rightarrow |
| | Pb(s) + KClO ₃ (aq.) X > |
| | 6Na(s) +2H3 POH (a) -> 2NO3POH (aq) + 3H2(g) |
| | |

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Sebastian Brünink

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