

Name Mel

Locker # _____

Bonding & Molecular Structure**LEWIS STRUCTURES**

Indicate the total # of VALENCE electrons for each molecule (VE's). Calculate the number of bonded electrons from the skeleton (BE's). Calculate the Remaining Electrons (RE's = VE's - BE's) and distribute the remaining electrons onto the skeletal structure. Finally, satisfy the octets by creating double and triple bonds. Your final structures should be given below and your scratch paper should clearly show your work.

F_2 VE's = <u>14</u> BE's = <u>2</u> RE's = <u>12</u> $\text{:}\ddot{\text{F}}-\ddot{\text{F}}\text{:}$ Shape: <u>linear</u>	H_2 VE's = <u>2</u> BE's = <u>2</u> RE's = <u>0</u> $\text{H}-\text{H}$ Shape: <u>linear</u>	HF VE's = <u>8</u> BE's = <u>2</u> RE's = <u>6</u> $\text{H}-\ddot{\text{F}}\text{:}$ Shape: <u>linear</u>	NH_3 VE's = <u>8</u> BE's = <u>6</u> RE's = <u>2</u> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{N}-\text{H} \\ \\ \cdot \end{array}$ Shape: <u>trigonal pyramidal</u>
CH_4 VE's = <u>8</u> BE's = <u>8</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$ Shape: <u>tetrahedral</u>	C_2H_6 VE's = <u>14</u> BE's = <u>14</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$ Shape: <u>tetrahedral (both)</u>	AlH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u> $\begin{array}{c} \text{H}-\text{Al}-\text{H} \\ \\ \text{H} \end{array}$ Shape: <u>trigonal planar</u>	BH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{B}-\text{H} \end{array}$ Shape: <u>trigonal planar</u>
C_2H_4 VE's = <u>12</u> BE's = <u>12</u> RE's = <u>0</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{C}=\text{C} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array}$ Shape around one C: <u>trigonal planar</u>	C_2F_4 VE's = <u>36</u> BE's = <u>12</u> RE's = <u>24</u> $\begin{array}{c} \text{:}\ddot{\text{F}} \quad \text{:}\ddot{\text{F}}\text{:} \\ \diagdown \quad \diagup \\ \text{C}=\text{C} \\ \diagup \quad \diagdown \\ \text{:}\ddot{\text{F}} \quad \text{:}\ddot{\text{F}}\text{:} \end{array}$ Shape around one C: <u>trigonal planar</u>	CO VE's = <u>10</u> BE's = <u>6</u> RE's = <u>4</u> $\text{:}\ddot{\text{C}}\equiv\ddot{\text{O}}\text{:}$ Shape: <u>linear</u>	O_2 VE's = <u>12</u> BE's = <u>4</u> RE's = <u>8</u> $\text{:}\ddot{\text{O}}=\ddot{\text{O}}\text{:}$ Shape: <u>linear</u>

South Pasadena Bonding And Molecular Structure Worksheet Answers

JA Banks



South Pasadena Bonding And Molecular Structure Worksheet Answers:

Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1980 *A Pictorial Approach to Molecular Bonding* John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made. It is therefore becoming increasingly important for the non-theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details associated with particular molecules. It seems highly desirable then that the non-theoretician be able to deduce results of general validity without esoteric mathematics. In this context pictorial reasoning is particularly useful. Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers. Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics. The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial. Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 . Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example.

Bonding and Molecular Structure Russell S. Drago, 1971 *Chemical Bonds* Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding. This is a reprint of a book first published by Benjamin Cummings 1973 *Structure of Molecules and the Chemical Bond* Y. K. Syrkin, M. E. Dyatkina, 1950

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