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A Solution Manual for: COMBUSTION, 4th Ed. by L Glassman, and R.A. Yetter CHAPTER I

Problem 1

Calculate the heat of reaction when methane and air in stoichiometric proportions are brought into a calorimeter at 500 K. The product composition is brought to the ambient temperature (298 K) by the cooling water. The pressure in the calorimeter is assumed to remain at 1 atm, but the water formed has condensed.

Solution: The heat of reaction can be calculated using Eqn. (10) in the text. However, as also suggested in the textbook, this equation can be simplified for the general case when $H_{\xi_i}^n = H_{2m}^n$. The modified equation becomes:

$$-Q = \Delta H = \sum_{i,j=0} n_i \left[\left(H^{\alpha}_{\tau_j} - H^{\alpha}_{2m} \right) + \left(\Delta H^{\alpha}_{\tau} \right)_{2m} \right]_i - \sum_{j,m=1} n_j \left[\left(H^{\alpha}_{\tau_j} - H^{\alpha}_{2m} \right)_{2m} + \left(\Delta H^{\alpha}_{\tau} \right)_{2m} \right]_j$$
(1)

In this Equation, T2 and T1 are the temperatures of the products and reactants respectively. The general equation for the complete combustion of one mole of methane in air is:

$$CH_a + a_{oc}(O_2 + 3.76N_2) \rightarrow aCO_2 + bH_2O + cN_2$$

where an is a coefficient describing how much air is needed for stoichiometric combustion. Balancing the above reaction one has:

$$CH_4 + 2(O_1 + 3.76N_1) \rightarrow CO_1 + 2H_2O + 7.52N_1$$

In order to apply Eqn. (1) to the above reaction, the following properties from Tables in the textbook are needed (note that for this problem $T_2 = 298$ K and $T_1 = 500$ K):

gases	ΔH ^o _c (kJ/mol)	$H_{T_i}^n - H_{pss}^n$ (kJ/mol)	H ⁰ _{T1} - H ⁰ ₂₀₀ (kJ/mol)
CH4	-74.90	8.21	0.0
O_2	0.0	6.09	0.0
N_2	0.0	5.92	0.0
H ₂ O _{Chi45}	-286.04	0.0	0.0
CO ₂	-393.77	0.0	0.0

Equation (1) then becomes:

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J Dewey

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