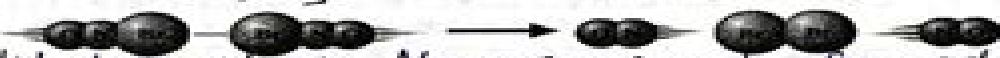


K/U:	17	TH:	10	C:	16	APP:	19
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1. Please answer the following questions in the space provided. [K5]

a) When solid KBr is dissolved in water, the solution gets colder. Is the chemical system exothermic or endothermic? [1]	endothermic
b) When a sample of potassium nitrate is dissolved in water, the temperature of the water changes from 23.5°C to 20.3°C. Is the chemical system exothermic or endothermic? [1]	endothermic
c) In the reaction $2A + B \rightarrow C + 3D$, reactant A is found to disappear at the rate of $6.2 \times 10^{-4} \text{ mol/L.s}$. What is the rate of disappearance of B? [1]	$\underline{2 \text{ mol of A disappear at rate } 6.2 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}}$ $\underline{1 \text{ mol of B disappears at } x \frac{\text{mol}}{\text{L.s}}}$ $x = \frac{6.2 \times 10^{-4} \text{ L.s}}{2} = 3.1 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$
d) The following diagram shows a collision between two BrNO molecules. Use collision theory to explain the results of the collision. [2]	 Reaction did take place. New compounds formed; 2 NO and 1 Br₂ are the products. It was an effective collision. Reactants collided with proper orientation and sufficient (optimal) kinetic energy.

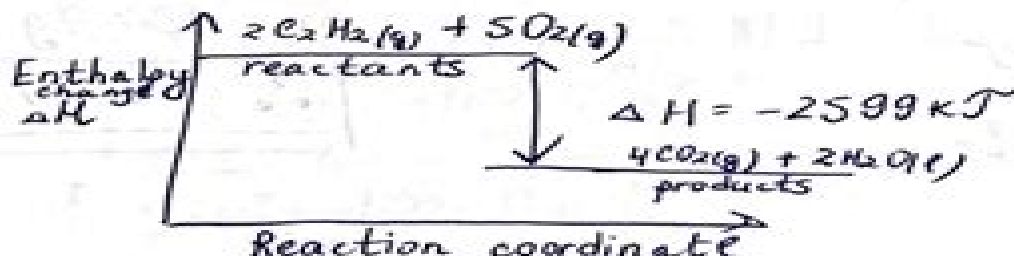
2. Consider the following reaction:



a) Write a thermochemical equation with the enthalpy change, ΔH , 'inside' the equation. [C1]



b) Sketch an enthalpy diagram for this reaction. Provide appropriate labels for each axis. [C3]



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Clifford Lane



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