## Chapter 18

#### Study Guide - Answers

- Write the expression for the equilibrium constant for each of the following reactions.
  - a.  $H_2(g) + Cl_2(g) = 2HCl(g)$

$$K_{eq} = \frac{[HCI]^2}{[H_1][CI_2]}$$

b. 
$$N_2(g) + 3 H_2(g) \cong 2 NH_3(g)$$

$$K_{eq} = \frac{[NH_0]^2}{[N_0][H_0]^2}$$

c. 
$$HCN(g) \cong H'(aq) + CN'(aq)$$

d. 
$$H_2SO_2(g) = H^*(aq) + HSO^*(aq)$$

$$K_{eq} = \frac{[H'][HSO^*]}{[HSO_0]}$$

e. 
$$PbF_2(s) = Pb^{2s}(aq) + 2F'(aq)$$

$$K_{eq} = - \|Pb^{2r}\||F'|^2$$

From the data provided below, calculate the value of the equilibrium constant for the reaction.

a. 
$$H_2(g) + Cl_2(g) = 2HCl(g)$$

$$[H_2] = [CI_2] = 1.0 \times 10^{-2}; [HCI] = 1.0 \times 10^{-1}$$

$$K_{eq} = \frac{[HCI]^2}{[H_2][CI_2]}$$

$$K_{eq} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-4}] \times [1.0 \times 10^{-4}]} = 0.0001$$

b. 
$$N_2(g) + 3 H_2(g) = 2 NH_2(g)$$

$$[N_2] = 4.4 \times 10^{-2}; [H_2] = 1.2 \times 10^{-1}; [NH_3] = 3.4 \times 10^{-3}$$

$$K_{eq} = \frac{[NH_1]^r}{[N_2][H_2]^r}$$

$$K_{eq} = \frac{[3.4 \times 10^{-4}]^{2}}{[4.4 \times 10^{-4}][1.2 \times 10^{-4}]^{2}} = 0.152$$

$$N_2O_1(g) = 2NO_2(g)$$

$$[N_2O_2] = 2.0 \text{ Ar}$$
  
 $[NO_2] = 0.2 \text{ Ar}$ 

$$Q = -\frac{[NO_2]^2}{[N_2O_2]}$$

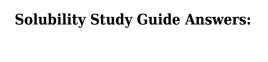
$$Q = \frac{\{0,2\}^2}{\{2,0\}} = 0.02$$

$$K_{\rm eq} = 0.2$$

 $Q < K_{\rm eq}$ , so the reaction will proceed to the right (forward).

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