

Des nuisances de la surface aux abysses

Que ce soit par les bateaux à la surface ou les robots dans les profondeurs, l'extraction des nodules polymétalliques menace les écosystèmes locaux.

Des effluents nocifs

Les effluents résiduels issus de la phase de séchage des minerais se composeraient principalement de sédiments et d'eau de mer, mais aussi de particules de minerais, et ils seraient rejetés dans l'eau par une conduite de rinjection. Or ce mélange pourrait intoxiquer la faune sous-marine.

Des panaches asphyxiants

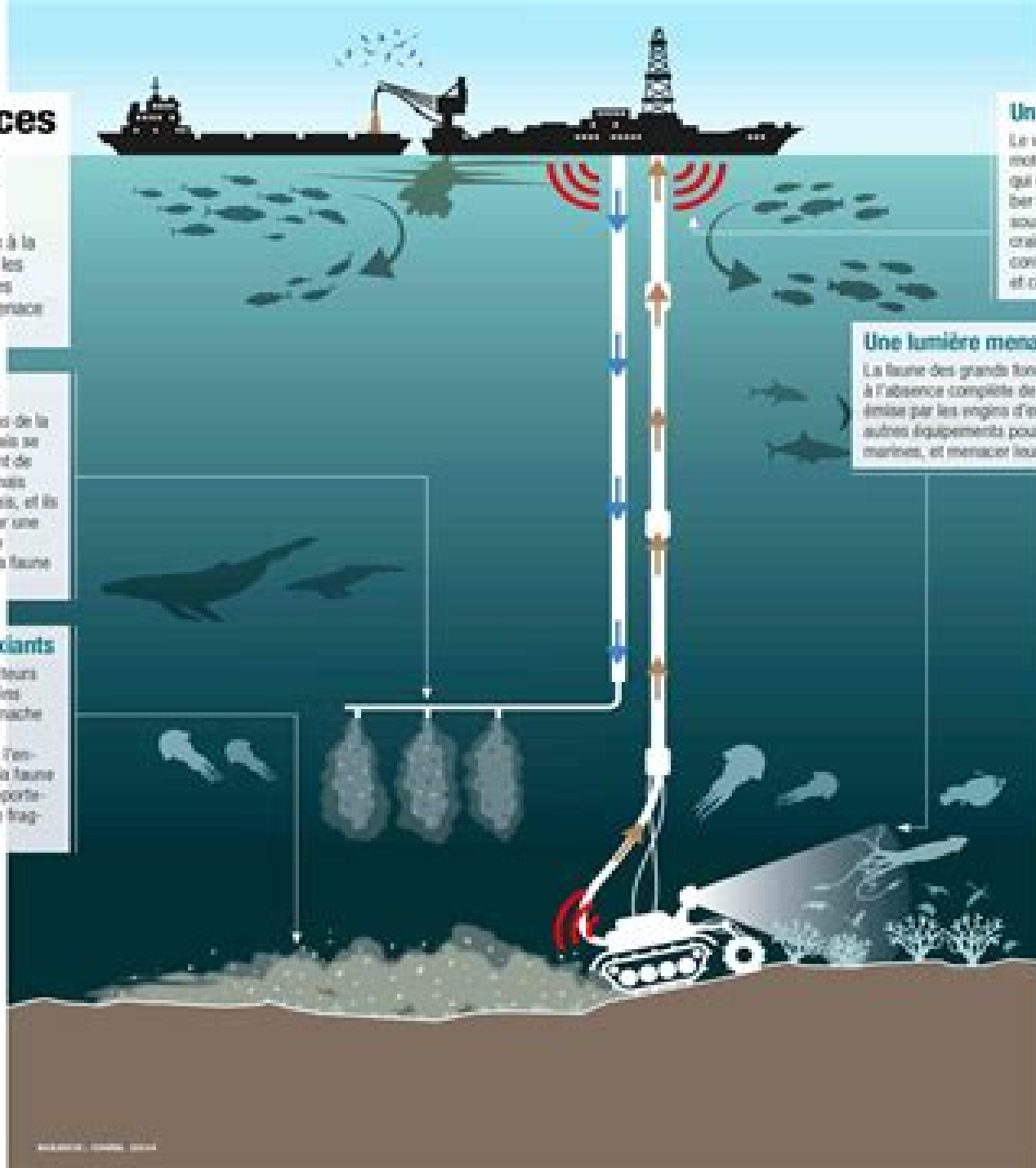
Le passage des robots collecteurs de nodules sur les fonds marins entraîne la formation d'un panache de sédiments. Ces panaches peuvent causer l'asphyxie ou l'entassement d'une partie de la faune et de la flore, modifier le comportement des espèces, ou encore fragmenter les habitats.

Un bruit perturbant

Le vacarme généré en surface par les moteurs, pompes et autres systèmes qui équipent les bateaux peut perturber le comportement des espèces sous-marines. Il existe également des craintes concernant de possibles conflits entre les activités d'extraction et celles de la pêche.

Une lumière menaçante

La faune des grands fonds est généralement habituée à l'absence complète de lumière. La pollution lumineuse émise par les engins d'extraction sous-marins et les autres équipements pourrait perturber les espèces sous-marines, et menacer leur santé et leur développement.



Socioanthropologie Des Marins Pecirccheurs

RJ Shavelson



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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no₂ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0ml of nitric acid and 4.0ml of sulfuric acid and will then react with 4.5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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