



## 2-2: Writing Balanced Precipitation Reactions **Lab chemistry**

In this problem, you will go into the virtual laboratory and perform a series of precipitation reactions using  $\text{Ag}^+$ ,  $\text{Pb}^{2+}$ , and  $\text{Sb}^{3+}$ . After observing the reactions, you will write the net ionic equations representing these reactions and then balance them.

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Writing Balanced Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of  $\text{Ag}^+$  ion solution (it is  $\text{AgNO}_3$ ) on the shelf to add it to the test tube. Then drag the test tube to the blue test tube rack. Repeat for  $\text{Pb}^{2+}$  ( $\text{Pb}(\text{NO}_3)_2$ ) and  $\text{Sb}^{3+}$  ( $\text{Sb}(\text{OH})_3$ ).
3. Moving your mouse over each test tube in the test-tube rack will allow you to identify it on the TV monitor. You can also click on the label at the top of each test tube in order to label the test tubes. Place the test tube containing the  $\text{Ag}^+$  solution in the metal test tube stand. You will be able to observe what happens in the window at the top left. Click on the  $\text{Na}_2\text{CO}_3$  reagent bottle to add it to the test tube in the stand.

*What color is the precipitate? Write out the full balanced molecular equation for the reaction.*

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*Write the balanced ionic equation for the reaction.*

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*Write the balanced net ionic equation for the reaction.*

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# Virtual Lab Writing Balanced Precipitation Reactions Answers

**ML Yell**



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