

## Ch 13 Study Guide (Chemical Bonds)

1. Fill in the table below with the correct values:

| Element Name | Ion Symbol | # Protons | # Electrons | Charge | Ion Type |
|--------------|------------|-----------|-------------|--------|----------|
| Fluorine     | $F^-$      | 9         | 10          | -1     | Anion    |
| Iodine       | $I^-$      | 53        | 54          | -1     | Anion    |
| Sulfur       | $S^{2-}$   | 16        | 18          | -2     | Anion    |
| Potassium    | $K^+$      | 19        | 18          | +1     | Cation   |
| Calcium      | $Ca^{2+}$  | 20        | 18          | +2     | Cation   |
| Bromine      | $Br^-$     | 35        | 36          | -1     | Anion    |
| Strontium    | $Sr^{2+}$  | 38        | 36          | +2     | Cation   |
| Oxygen       | $O^{2-}$   | 8         | 10          | -2     | Anion    |
| Magnesium    | $Mg^{2+}$  | 12        | 10          | +2     | Cation   |
| Aluminum     | $Al^{3+}$  | 13        | 10          | +3     | Cation   |
| Selenium     | $Se^{2-}$  | 34        | 36          | -2     | Anion    |
| Lithium      | $Li^+$     | 3         | 2           | +1     | Cation   |
| Rubidium     | $Rb^+$     | 37        | 36          | +1     | Cation   |
| Chlorine     | $Cl^-$     | 17        | 18          | -1     | Anion    |

2. An ionic bond is formed between what two types of atoms?

Metal & nonmetal

3. A covalent bond is formed between what two types of atoms?

Nonmetal & Nonmetal

4. Draw the bond(s) that will form between Mg and O.



5. What is the formula for the compound formed in question #4?



# Simple Ions Answer Key

**Robert Boxer**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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