

Chemistry Spring Final Study Guide

Unit 10, Acids & Bases
Fig. 10.10: Rank the acids and compare the properties of acids. All bases.

Acids	Bases
1. Hydrochloric acid	1. Sodium hydroxide
2. Sulfuric acid	2. Potassium hydroxide
3. Nitric acid	3. Calcium hydroxide
4. Acetic acid	4. Ammonia
5. Phosphoric acid	5. Sodium carbonate

Unit 11, Solubility
Classify the following compounds as soluble or insoluble in water. Write the formula of the precipitate if a precipitate forms.

1. NaCl (soluble)
2. AgCl (insoluble)
3. KNO_3 (soluble)
4. BaSO_4 (insoluble)
5. Fe(OH)_3 (insoluble)
6. PbCl_2 (insoluble)
7. CaCO_3 (insoluble)
8. Mg(OH)_2 (insoluble)
9. LiCl (soluble)
10. CuS (insoluble)

Unit 12, Thermodynamics
Calculate the enthalpy change for the following reactions.

1. $\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$
2. $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$
3. $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
4. $\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$
5. $\text{Fe}_2\text{O}_3(s) + 3\text{CO}(g) \rightarrow 2\text{Fe}(s) + 3\text{CO}_2(g)$

Unit 13, Equilibrium
Calculate the equilibrium constant for the following reactions.

1. $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$
2. $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$
3. $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$
4. $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$
5. $\text{Fe}_3\text{O}_4(s) + 4\text{H}_2(g) \rightleftharpoons 3\text{Fe}(s) + 4\text{H}_2\text{O}(g)$

Unit 14, Organic Chemistry
Identify the functional groups in the following molecules.

1. $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ (alcohol)
2. $\text{CH}_3\text{CH}_2\text{CHO}$ (aldehyde)
3. $\text{CH}_3\text{CH}_2\text{COCH}_3$ (ketone)
4. $\text{CH}_3\text{CH}_2\text{NH}_2$ (amine)
5. $\text{CH}_3\text{CH}_2\text{COOH}$ (carboxylic acid)

Unit 15, Biochemistry
Identify the biomolecules in the following list.

1. Glucose (carbohydrate)
2. Amino acids (proteins)
3. Nucleic acids (DNA/RNA)
4. Lipids (fats/oils)
5. Enzymes (proteins)

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Unit 1, Atomic Structure
Rank the following elements from most reactive to least reactive.

1. Sodium (Na)
2. Potassium (K)
3. Calcium (Ca)
4. Magnesium (Mg)
5. Aluminum (Al)
6. Silicon (Si)
7. Phosphorus (P)
8. Sulfur (S)
9. Chlorine (Cl)
10. Bromine (Br)
11. Iodine (I)
12. Fluorine (F)
13. Oxygen (O)
14. Nitrogen (N)
15. Carbon (C)
16. Boron (B)
17. Beryllium (Be)
18. Lithium (Li)
19. Hydrogen (H)
20. Helium (He)

Unit 2, Periodic Trends
Rank the following elements from most electronegative to least electronegative.

1. Fluorine (F)
2. Oxygen (O)
3. Nitrogen (N)
4. Carbon (C)
5. Silicon (Si)
6. Aluminum (Al)
7. Magnesium (Mg)
8. Calcium (Ca)
9. Potassium (K)
10. Sodium (Na)

Unit 3, Bonding
Rank the following compounds from most ionic to least ionic.

1. NaCl
2. KBr
3. CaF_2
4. MgO
5. Al_2O_3
6. SiO_2
7. CO_2
8. CH_4
9. H_2O
10. NH_3

Unit 4, Solutions
Rank the following solutions from most concentrated to least concentrated.

1. 10% NaCl solution
2. 5% NaCl solution
3. 1% NaCl solution
4. 0.1% NaCl solution
5. 0.01% NaCl solution

Unit 5, Acids & Bases
Rank the following acids from most acidic to least acidic.

1. HCl
2. H_2SO_4
3. HNO_3
4. H_2CO_3
5. CH_3COOH
6. H_2O
7. NH_3
8. NaOH
9. KOH
10. Ca(OH)_2

Unit 6, Solubility
Rank the following compounds from most soluble to least soluble in water.

1. NaCl
2. KNO_3
3. CaSO_4
4. BaSO_4
5. Fe(OH)_3
6. PbCl_2
7. CuS
8. Mg(OH)_2
9. LiCl
10. AgCl

Unit 7, Thermodynamics
Rank the following reactions from most exothermic to least exothermic.

1. $\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$
2. $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$
3. $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
4. $\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$
5. $\text{Fe}_2\text{O}_3(s) + 3\text{CO}(g) \rightarrow 2\text{Fe}(s) + 3\text{CO}_2(g)$

Unit 8, Equilibrium
Rank the following reactions from most spontaneous to least spontaneous.

1. $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
2. $\text{H}_2(g) + \text{I}_2(g) \rightarrow 2\text{HI}(g)$
3. $\text{PCl}_5(g) \rightarrow \text{PCl}_3(g) + \text{Cl}_2(g)$
4. $\text{CO}_2(g) + \text{H}_2(g) \rightarrow \text{CO}(g) + \text{H}_2\text{O}(g)$
5. $\text{Fe}_3\text{O}_4(s) + 4\text{H}_2(g) \rightarrow 3\text{Fe}(s) + 4\text{H}_2\text{O}(g)$

Unit 9, Organic Chemistry
Rank the following molecules from most polar to least polar.

1. $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
2. $\text{CH}_3\text{CH}_2\text{CHO}$
3. $\text{CH}_3\text{CH}_2\text{COCH}_3$
4. $\text{CH}_3\text{CH}_2\text{NH}_2$
5. $\text{CH}_3\text{CH}_2\text{COOH}$

Unit 10, Biochemistry
Rank the following biomolecules from most abundant to least abundant.

1. Glucose
2. Amino acids
3. Nucleic acids
4. Lipids
5. Enzymes

Spring Semester Environmental Chemistry Final Study Guide

Harold R. Doughty



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Resources in Education, 1997 EC Study Guide to Environment-related Courses Institute for European Environmental Policy, Commission of the European Communities. Directorate-General Environment, Nuclear Safety and Civil Protection, 1993 **Annual Catalogue** United States Air Force Academy, 1984 **Guide to American Graduate Schools**

Harold R. Doughty, 2009-02-24 For students planning further study after college the Guide to American Graduate Schools puts the necessary information at their fingertips Completely revised and updated this long trusted and indispensable tool features comprehensive information on every aspect of graduate and professional study including Alphabetically arranged profiles of more than 1 200 accredited institutions including enrollment locations libraries and other facilities and housing situations Fields of study offered by each institution and types of degrees conferred Admissions standards and requirements recruitment practices and degree requirements Tuition costs and opportunities for financial aid Details on scholarships fellowships assistantships and internships Organized in a clear straightforward easy to use format this is the essential source with which to begin planning for the future United States Air Force Academy United States Air Force Academy,

Bulletin of Information United States Coast Guard Academy, Broadening Participation in STEM Zayika Wilson-Kennedy, Goldie S. Byrd, Eugene Kennedy, Henry T. Frierson, 2019-02-28 This book reports on high impact educational practices and programs that have been demonstrated to be effective at broadening the participation of underrepresented groups in the STEM disciplines *Research in Education*, 1974 University Record University of Chicago, 1903

Undergraduate Catalog ... with Announcements for ... Western Michigan University, 1929 *University Record*, 1903

Environmental Organic Chemistry for Engineers James G. Speight, 2016-11-02 Environmental Organic Chemistry for Engineers clearly defines the principles of environmental organic chemistry and the role they play in forming remediation strategies In this reference the author explores parameter estimation methods the thermodynamics and kinetics needed to predict the fate transports and reactivity of organic compounds in air water and soils The book s four part treatment starts with the classification of organic molecules and physical properties of natural organic matter halocarbons phenols polyaromatic hydrocarbons organophosphates and surfactants An overview of remediation technologies and a discussion of the interactions that lead to physical properties that affect chemical distribution in the environment is also detailed as are the important reaction classes of organic molecules including substituent effects and structure and activity relationships found in Part Two and Three Part four is devoted to the strengths and weaknesses of different remediation technologies and when they should be employed Clearly defines the principles of environmental organic chemistry and the role they play in forming remediation strategies Includes the tools and methods for classifying environmental contaminants found in air water and soil Presents a wide range of remediation technologies and when they should be deployed for maximum effect How Finns Learn Mathematics and Science, 2007-01-01 The Finnish students success in the first PISA 2000 evaluation was a

surprise to most of the Finns and even people working in teacher education and educational administration had difficulties to believe that this situation would continue Finland's second success in the next PISA 2003 comparison has been very pleasing for teachers and teacher educators and for education policymakers The good results on the second time waked us to think seriously on possible reasons for the success Several international journalists and expert delegations from different countries have asked these reasons while visiting in Finland Since we had no commonly acceptable explanation to students success we decided at the University of Helsinki to put together a book *How Finns Learn Mathematics and Science* in order to give a commonly acceptable explanation to our students success in the international PISA evaluations The book tries to explain the Finnish teacher education and school system as well as Finnish children's learning environment at the level of the comprehensive school and thus give explanations for the Finnish PISA success The book is a joint enterprise of Finnish teacher educators The explanations for success given by altogether 40 authors can be classified into three groups Teacher and teacher education school and curriculum and other factors like the use of ICT and a developmental project LUMA The main result is that there is not one clear explanation although research based teacher education seems to have some influence But the true explanation may be a combination of several factors

Basics of Ecotoxicology Donald W. Sparling, 2017-07-28 This textbook presents a comprehensive examination of environmental science and ecotoxicology for undergraduate students The material provides sufficient related background information leading to a competency to clearly understand ecotoxicology concepts and topics

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