

AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$ (notice the d-block)
 - Nickel (with shorthand): $[\text{Ar}] 4s^2 3d^8$
- **Law of Definite Proportions:**
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

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Kaplan

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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scored 65% and 55% on his first two exams and 95% on his third. These are representative of thousands of students who decisively improved their grades by acting on the advice described in this book. What is preventing your students from performing according to expectations? Sandra McGuire offers a simple but profound answer: If you teach students how to learn and give them simple straightforward strategies to use, they can significantly increase their learning and performance. For over a decade, Sandra McGuire has been acclaimed for her presentations and workshops on metacognition and student learning because the tools and strategies she shares have enabled faculty to facilitate dramatic improvements in student learning and success. This book encapsulates the model and ideas she has developed in the past fifteen years—ideas that are being adopted by an increasing number of faculty with considerable effect. The methods she proposes do not require restructuring courses or an inordinate amount of time to teach. They can often be accomplished in a single session, transforming students from memorizers and regurgitators to students who begin to think critically and take responsibility for their own learning. Sandra McGuire takes the reader sequentially through the ideas and strategies that students need to understand and implement. First, she demonstrates how introducing students to metacognition and Bloom's Taxonomy reveals to them the importance of understanding how they learn and provides the lens through which they can view learning activities and measure their intellectual growth. Next, she presents a specific study system that can quickly empower students to maximize their learning. Then, she addresses the importance of dealing with emotion, attitudes, and motivation by suggesting ways to change students' mindsets about ability and by providing a range of strategies to boost motivation and learning. Finally, she offers guidance to faculty on partnering with campus learning centers. She pays particular attention to academically unprepared students, noting that the strategies she offers for this particular population are equally beneficial for all students. While stressing that there are many ways to teach effectively and that readers can be flexible in picking and choosing among the strategies she presents, Sandra McGuire offers the reader a step-by-step process for delivering the key messages of the book to students in as little as 50 minutes. Free online supplements provide three slide sets and a sample video lecture. This book is written primarily for faculty but will be equally useful for TAs, tutors, and learning center professionals. For readers with no background in education or cognitive psychology, the book avoids jargon and esoteric theory.

Organic Chemistry Education Research into Practice Jay Wackerly, Sarah Zingales, Michael Wentzel, Gautam Bhattacharyya, Brett McCollum, 2025-03-25. This Research Topic has three main goals: 1) provide a platform for instructors of organic chemistry to showcase evidence-based methods and educational theories they have utilized in their classrooms; 2) build new and strengthen existing connections between educational researchers and practitioners; and 3) highlight how people have used chemical education-based research in their teaching practice. There are places in the literature dedicated for chemical education research (CER); however, there is not a clear avenue for those that have changed their teaching methods based on published CER and report their experiences. Creating this article collection will foster collaboration between

chemical education researchers and teachers of organic chemistry This opportunity allows these instructors to share evidence based practices experiences challenges and innovative approaches from CER literature and beyond This Research Topic bridges discipline based education research and the scholarship of teaching and learning which will help advance organic chemistry education and improve student outcomes

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