

Annotated periodic table of the elements

1	2											13	14	15	16	17	18		
1 H 1.01 Hydrogen												3 Li 6.94 Lithium	4 Be 9.01 Beryllium	5 B 10.81 Boron	6 C 12.01 Carbon	7 N 14.01 Nitrogen	8 O 16.00 Oxygen	9 F 18.99 Fluorine	10 Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.01 Beryllium	Most common oxidation states: electronegativity (Mulliken-Pauling)										13 Al 26.98 Aluminium	14 Si 28.09 Silicon	15 P 30.97 Phosphorus	16 S 32.06 Sulfur	17 Cl 35.45 Chlorine	18 Ar 39.95 Argon		
11 Na 22.99 Sodium	12 Mg 24.31 Magnesium	Group name* 1 Hydrogen & the alkali metals 2 Alkaline earth metals 13 Coinage metals 14 Volatile metals 15 Isotopes 16 Crystallogens 17 Pnictogens 18 Chalcogens 19 Halogens 20 Halium group *Groups 1 to 10 are named after their first members i.e. Group 1 is the Scandium group										31 Ga 69.72 Gallium	32 Ge 72.64 Germanium	33 As 74.92 Arsenic	34 Se 78.96 Selenium	35 Br 79.90 Bromine	36 Kr 83.80 Krypton		
19 K 39.10 Potassium	20 Ca 40.08 Calcium	21 Sc 44.96 Scandium	22 Ti 47.87 Titanium	23 V 50.94 Vanadium	24 Cr 52.00 Chromium	25 Mn 54.94 Manganese	26 Fe 55.85 Iron	27 Co 58.93 Cobalt	28 Ni 58.69 Nickel	29 Cu 63.55 Copper	30 Zn 65.38 Zinc	31 Ga 69.72 Gallium	32 Ge 72.64 Germanium	33 As 74.92 Arsenic	34 Se 78.96 Selenium	35 Br 79.90 Bromine	36 Kr 83.80 Krypton		
37 Rb 85.47 Rubidium	38 Sr 87.62 Strontium	39 Y 88.91 Yttrium	40 Zr 91.22 Zirconium	41 Nb 92.91 Niobium	42 Mo 95.94 Molybdenum	43 Tc 98.91 Technetium	44 Ru 101.07 Ruthenium	45 Rh 101.07 Rhodium	46 Pd 106.32 Palladium	47 Ag 107.87 Silver	48 Cd 112.41 Cadmium	49 In 114.82 Indium	50 Sn 118.71 Tin	51 Sb 121.76 Antimony	52 Te 127.60 Tellurium	53 I 126.91 Iodine	54 Xe 131.29 Xenon		
55 Cs 132.91 Cesium	56 Ba 137.33 Barium	57 to 70										81 Tl 204.38 Thallium	82 Pb 207.2 Lead	83 Bi 208.98 Bismuth	84 Po 209 Polonium	85 At 210 Astatine	86 Rn 222 Radon		
87 Fr 223 Francium	88 Ra 226 Radium	89 to 102										113 Nh 286.10 Nihonium	114 Fl 289.10 Flerovium	115 Mc 288.10 Moscovium	116 Lv 289.10 Livermorium	117 Ts 284.10 Tennessine	118 Og 294.10 Oganesson		

- Group names*
- Hydrogen & the alkali metals
 - Alkaline earth metals
 - Coinage metals
 - Volatile metals
 - Isotopes
 - Chalcogens
 - Chalcogens
 - Halogens
 - Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
(Alred-Pauling)

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals. Those from Th onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metals

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- Group 1**
The alkali metals (Li, Na, K, Rb, Cs, Fr) and the alkaline earth metals (Be, Mg, Ca, Sr, Ba, Ra) are the most reactive metals in the periodic table.

Nonmetals

- Chalcogens**
O, S, Se, and Te are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Unscrambling The Periodic Table Answers

Chris McMullen, Carolyn Kivett



chemical words A chemical word is a word that can be made using symbols from the periodic table For example the chemical word POSSEsSiON is made using the symbols for phosphorus P oxygen O sulfur S Einsteinium Es silicon Si and nitrogen N You don t need to be familiar with the periodic table to solve these problems nor do you need to know any chemistry You just need to be able to count and unscramble elements to make words This Easy volume consists of words with 4 to 5 symbols which involves familiarity with common 4 to 10 letter words Other Medium and Hard volumes consist of longer words A unique feature of this book is that there is a Hints section at the back separate from the Answers section for puzzlers who may be stuck and want to check just the first letter of the solution MORE EXAMPLES 1 S Ni Ge U _ _ _ _ 2 2 C N 2 I P _ _ _ _ 3 Ti C Cr P Y _ _ _ _ 4 2 C U 2 S Es _ _ _ _ You can find the answers below Note that this easy volume consists of chemical words with 4 to 5 symbols We recommend starting with our easy puzzles before tackling the medium or hard puzzles available in separate volumes ANSWERS 1 GeNiUS 2 PICNIC 3 CrYPTiC 4 SUCCEsS

Jumpstarters for Properties of Matter, Grades 4 - 12 Eric T. Olson, 2008-12-19 Engage students in describing and comparing general properties of objects like size shape and temperature Then move on to specific types of matter that have characteristic properties such as density and the melting point And lastly have students focus on the notions of elements and compounds Each reproducible page includes five exercises that can be used as whole worksheets for homework assignments and extra practice or cut apart for daily warm up activities Supports NSE standards *Amazing People: Scientists and Mathematicians* Elise Craver, Malaski, Malay, Hailey Scragg, 2023-01-23 Inspirational 1st 2nd and 3rd grade workbooks for kids ages 7 Motivate and inspire children with the Amazing People Scientists and Mathematicians Activity Book 1st 2nd and 3rd grade books are a great way for children to learn about inspirational people who have made important contributions to science and math through a variety of activities that are both fun AND educational Why You ll Love This 1st Grade Third Grade Workbook Engaging and educational activities and wall d cor The activity book features biographies activities and flash cards covering 53 inspirational mathematicians and scientists The activity book also includes 1 inspirational poster great for both homeschool and classroom curriculum and classroom d cor Tracking progress along the way Test your child s knowledge with the flash cards that cover scientists and mathematicians learned throughout the book Use the stickers to reward students on a job well done after completing activities Practically sized for every activity The 256 page 1st grade workbook is sized at 7 75 x 10 625 giving your child plenty of space to complete each exercise About Carson Dellosa For more than 45 years Carson Dellosa has provided solutions for parents and teachers to help their children get ahead and exceed learning goals Carson Dellosa supports your child s educational journey every step of the way The Grade 1 Grade 3 Workbook Contains Biographies and activities covering 53 people 64 flash cards 53 stickers 1 motivational poster

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Verbal Reactions - Word Scrambles With a Chemical Flavor (Hard) Chris

Mcmullen,Carolyn Kivett,2011-06-18 VErBAL ReAcTiONS are a new form of word scrambles with a chemical flavor You don t need to know any science to be able to solve VErBAL ReAcTiONS puzzles but the puzzles bear a resemblance to chemical reactions Here is a sample Es 2 S P Si 2 O N _ _ _ _ _ This VErBAL ReAcTiON is a word scramble consisting of one Es two S s one P one Si two O s and one N That is the word scramble contains the elements Es S S P Si O O and N Unscramble these elements to form an 8 symbol word that s why there are 8 blanks in the puzzle For this puzzle the answer is P O S S Es Si O N possession These VErBAL ReAcTiONS resemble chemical reactions in two ways First the scrambled elements appear added together on the left of the reaction with coefficients like the number 2 in the puzzle above telling you how many of each element the solution contains and you fill in the result of the VErBAL ReAcTiON by rearranging the elements and writing them on the blanks on the right side of the reaction Secondly all of the solutions are chemical words A chemical word is a word that can be made using symbols from the periodic table For example the chemical word POSSEsSiON is made using the symbols for phosphorus P oxygen O sulfur S Einsteinium Es silicon Si and nitrogen N You don t need to be familiar with the periodic table to solve these problems nor do you need to know any chemistry You just need to be able to count and unscramble elements to make words This Hard volume consists of words with 7 to 8 symbols which involves familiarity with common 8 to 14 letter words Other Medium and Easy volumes consist of shorter words A unique feature of this book is that there is a Hints section at the back separate from the Answers section for puzzlers who may be stuck and want to check just the first letter of the solution MORE EXAMPLES 1 S Ni Ge U _ _ _ _ _ 2 2 C N 2 I P _ _ _ _ _ 3 Ti C Cr P Y _ _ _ _ _ 4 2 C U 2 S Es _ _ _ _ _ You can find the answers at the end of this paragraph Note that this hard volume consists of chemical words with 7 to 8 symbols which are longer than the examples shown here We recommend starting with our easy or medium puzzles before tackling these hard puzzles available in separate volumes ANSWERS 1 GeNiUS 2 PICNIC 3 CrYPTiC 4 SUCCEsS

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Science News-letter ,1962-07

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Table of Contents Unscrambling The Periodic Table Answers

1. Understanding the eBook Unscrambling The Periodic Table Answers
 - The Rise of Digital Reading Unscrambling The Periodic Table Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unscrambling The Periodic Table Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unscrambling The Periodic Table Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unscrambling The Periodic Table Answers
 - Personalized Recommendations
 - Unscrambling The Periodic Table Answers User Reviews and Ratings
 - Unscrambling The Periodic Table Answers and Bestseller Lists

5. Accessing Unscrambling The Periodic Table Answers Free and Paid eBooks
 - Unscrambling The Periodic Table Answers Public Domain eBooks
 - Unscrambling The Periodic Table Answers eBook Subscription Services
 - Unscrambling The Periodic Table Answers Budget-Friendly Options
6. Navigating Unscrambling The Periodic Table Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unscrambling The Periodic Table Answers Compatibility with Devices
 - Unscrambling The Periodic Table Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unscrambling The Periodic Table Answers
 - Highlighting and Note-Taking Unscrambling The Periodic Table Answers
 - Interactive Elements Unscrambling The Periodic Table Answers
8. Staying Engaged with Unscrambling The Periodic Table Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unscrambling The Periodic Table Answers
9. Balancing eBooks and Physical Books Unscrambling The Periodic Table Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unscrambling The Periodic Table Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unscrambling The Periodic Table Answers
 - Setting Reading Goals Unscrambling The Periodic Table Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unscrambling The Periodic Table Answers
 - Fact-Checking eBook Content of Unscrambling The Periodic Table Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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