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 Campus ID: 991-22-8629
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 Units: ENGL 4 MATH 4 BUS 4
 Broward Community College
 Business Undergraduate / Finance
 Business Undergraduate / International Business Certificate
 Business Undergraduate / Information Technology Management

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 Advisor(s): James Hardy
 Rachel Davis

Program: Evaluation Date: DEC-2010
 Fort Lauderdale FL

---Academic History---

Program	Business Undergraduate	Fall 2014	UWM Credit					
2014-05-15	Business Undergrad - Intended	ART 100	American Art and Culture (2)	DECA 3.00	A-	11.01		
2015-05-21	Finance	ECOR 103	Principles of Macroeconomics	DECA 3.00	A	12.00		
2015-05-21	Marketing	ECOR 104	Principles of Macroeconomics	DECA 3.00	A	12.00		
2016-01-08	Finance	ASTRON 103	Survey of Astronomy	DECA 3.00	A	12.00		
2016-01-08	Information Technology Management	PSYCH 101	Introduction to Psychology	DECA 3.00	A	12.00		
2016-06-21	Finance	MATH 105	Intermediate Algebra	3.00	A	12.00		
2016-06-21	International Business Certificate	Sum GPA: 3.945	GPA Credit: 18.00	Progress Credit: 18.00		71.01		
2016-06-21	Information Technology Management	Cum GPA: 3.945	Cum Totals: 18.00	Cum Totals(Degree): 21.00		71.01		
WINT 2011	Transfer Credit(s)GRD	2015-01-05	Dean's Honor List					
ASH 2010	History	3 B HIST 101	GERM 3.00					
BUS 2041	Law I	3 A BUS ADM 201	3.00					
ENC 1101	Composition I	3 C ENGLISH 101	3.00					
MAT 1023	Inter Algebra	3 C MATH 95	0.00					
PHI 1100	Intro Logic	3 B PHILOS 211	GERM 3.00					
From: Broward Community College			12.00					
SUMR 2011	Transfer Credit(s)GRD	2015-01-05	Dean's Honor List					
ACC 2001	Princ Acct I	3 B BUS ADM X	3.00					
MAT 1105	Calc Algebra	3 C MATH 95	0.00					
From: Broward Community College			3.00					
FALL 2011	Transfer Credit(s)GRD	2015-01-05	Dean's Honor List					
ACC 2011	Princ Acct II	3 B BUS ADM 201	3.00					
BUS 2042	Bus Law II	3 B BUS ADM X	3.00					
ENC 1102	Composition II	3 B ENGLISH 102	3.00					
ELM 1000	Survey West Civ I	3 C HIST 101	GERM 3.00					
SPC 1658	Pub Speaking	3 B COMMUN 103	GERM 3.00					
STA 2023	Statistics	3 D MATHSTAT 215	GERM 3.00					
From: Broward Community College			18.00					
Spring 2015		2015-05-26	Dean's Honor List					
BUS ADM 200	Business Writing	3.00	B+	9.00				
BUS ADM 200	Principles of Finance	3.00	A	12.00				
BUS ADM 210	Intro to Management Statistics	4.00	A	16.00				
ASTRON 104	Astronomy Laboratory	GERM 1.00	A	4.00				
MATH 211	Survey Calc/Analytic Geometry	GERM 4.00	A	16.00				
ENGLISH 205	Business Writing	3.00	B+	9.00				
Sum GPA: 3.777	GPA Credit: 18.00	Progress Credit: 18.00		67.00				
Cum GPA: 3.861	Cum Totals: 36.00	Cum Totals(Degree): 69.00		134.00				
Summer 2015		2015-05-26	Dean's Honor List					
BUS ADM 302	Managerial Accounting	3.00	A	12.00				
BUS ADM 310	Organizations	3.00	A	12.00				
BUS ADM 340	Principles of Marketing	3.00	A	12.00				
Sum GPA: 4.000	GPA Credit: 9.00	Progress Credit: 9.00		36.00				
Cum GPA: 3.699	Cum Totals: 45.00	Cum Totals(Degree): 78.00		174.00				
2015-08-26		No Action						

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the objectives of the boiling point determination lab were to determine the boiling point of a mystery substance provided to each student boiling point is the point at which a liquid changes physical states and becomes a vapor or gas boiling point is important to understand because it can be used to help distinguish between different substances

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at agiven pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube end c

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boiling point determination lab report conclusion objectivesto correctly measure boiling issues of various liquids and identify

an unknown liquid based on its boiling point background information objekte can be distinguished into acid properties and physical properties

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown compounds boiling points are affect by three key factors the first being intermolecular forces inter molecular affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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clean up solid and liquid spills immediately true g conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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nov 27 2019 essay sample the purpose of this experiment is to determine the boiling points of various organic compounds and to use these to identify unknowns equipment free essays

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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aug 11 2020 boiling points are very sensitive to changes in applied pressure so all boiling points should be reported with the measured pressure a compound s normal boiling point refers to its boiling point at a pressure of 760 text mm ce hg

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given pressure to determine the vapour pressure of a pure liquid at various temperatures

exploring boiling points sharmin khatun chem 1141 studocu - May 17 2023

exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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laboratory partners samantha and mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that needs to be determined by physical properties such as boiling point melting point density solubility etc in this experiment I am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a Bunsen burner boiling

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there are a variety of methods by which a sample's boiling point can be determined including distillation reflux and by using a Thiele tube the most straightforward method uses a Thiele tube and has the advantage of using less than 0.5 mL of material

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