

When stoichiometric equivalent moles of acid and base react, they are said to be at the *equivalence point*. The process of reaching the equivalence point with acids and bases is called *neutralization*.

In order to know when the equivalence point is reached, an indicator is added to the titration vessel. An *indicator* is a compound that will react with one of the titration species and for which a color change is observed very near the equivalence point, when the appropriate indicator has been selected. The indicator used in this experiment is called phenolphthalein which turns from clear in acid solution to a pink color when excess base is present.

Sample calculation

If 0.558g of potassium hydrogen phthalate (KHP) is dissolved in 50 mL of deionized water in a flask, what is the concentration of a NaOH solution if 12.78 mL are required to titrate the potassium hydrogen phthalate to the equivalence point?

- a) Determine the moles of potassium hydrogen phthalate (KHP) in the flask using mass and molar mass:

$$\text{moles KHP} = \text{mass KHP} \times \frac{1}{M_{\text{KHP}}} = 0.558\text{g} \times \frac{1 \text{ mol}_{\text{KHP}}}{204.23\text{g}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{KHP}}$$

- b) Using the stoichiometric relationship in the balanced chemical reaction (Eq.1), determine the moles of base required to reach the equivalence point:

$$\text{mole NaOH} = \text{mol}_{\text{KHP}} \times \frac{1 \text{ mol}_{\text{NaOH}}}{1 \text{ mol}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{KHP}} \times \frac{1 \text{ mol}_{\text{NaOH}}}{1 \text{ mol}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{NaOH}}$$

- c) Determine the concentration of NaOH by dividing the number of mole by volume of NaOH solution added:

$$\text{molarity of NaOH} = \frac{0.00273 \text{ mol}_{\text{NaOH}}}{0.01278 \text{ L}_{\text{NaOH solution}}} = 0.214 \text{ M NaOH}$$

After the $\approx 0.3 \text{ M}$ NaOH solution has been standardized (its accurate average molarity determined by titration), it can be referred to as a standardized solution and its average molarity from the three titrations should be written on the bottle. It, in turn, will be employed to determine the acetic acid content in household vinegar. Like potassium hydrogen phthalate, acetic acid (ethanoic acid) has one acidic proton, H^+ . Because the acid is monoprotic, it will react one-to-one with NaOH (Eq. 2).



This time you will be starting with a dilute solution of acetic acid in the flask instead of a solid, and you will know the accurate molarity of the NaOH solution (the average of the three values determined). After the titration, knowing the volume of NaOH used in the titration, you can determine the mole of NaOH needed to reach the equivalence point with the acetic acid:

$$\text{moles NaOH} = M_{\text{NaOH}} \left(\frac{\text{mol}}{\text{L}} \right) \times V_{\text{NaOH}} (\text{L})$$

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Table of Contents Uga Chem 1211 Lab Manual

1. Understanding the eBook Uga Chem 1211 Lab Manual
 - The Rise of Digital Reading Uga Chem 1211 Lab Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Uga Chem 1211 Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Uga Chem 1211 Lab Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Uga Chem 1211 Lab Manual
 - Personalized Recommendations
 - Uga Chem 1211 Lab Manual User Reviews and Ratings
 - Uga Chem 1211 Lab Manual and Bestseller Lists
5. Accessing Uga Chem 1211 Lab Manual Free and Paid eBooks
 - Uga Chem 1211 Lab Manual Public Domain eBooks
 - Uga Chem 1211 Lab Manual eBook Subscription Services
 - Uga Chem 1211 Lab Manual Budget-Friendly Options
6. Navigating Uga Chem 1211 Lab Manual eBook Formats

- ePub, PDF, MOBI, and More
- Uga Chem 1211 Lab Manual Compatibility with Devices
- Uga Chem 1211 Lab Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Uga Chem 1211 Lab Manual
 - Highlighting and Note-Taking Uga Chem 1211 Lab Manual
 - Interactive Elements Uga Chem 1211 Lab Manual
- 8. Staying Engaged with Uga Chem 1211 Lab Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Uga Chem 1211 Lab Manual
- 9. Balancing eBooks and Physical Books Uga Chem 1211 Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Uga Chem 1211 Lab Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Uga Chem 1211 Lab Manual
 - Setting Reading Goals Uga Chem 1211 Lab Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Uga Chem 1211 Lab Manual
 - Fact-Checking eBook Content of Uga Chem 1211 Lab Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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