

Activity A:**Molar Mass**Get the Gizmo ready:

- Select the AVOGADRO CONSTANT tab.
- Turn on **Show hints** and check that **Copper (Cu)** is selected.



Introduction: Since atoms are so tiny, chemists have devised a unit known as the **mole**. A mole represents a macroscopic quantity of matter that can be used in the laboratory. One mole of any element has the same mass in grams as its **atomic mass** in u.

Question: How many particles are in a mole?

1. **Explore:** Note the average atomic mass of copper on the nano-balance. Add atoms to the larger balance until it registers the same number (in g) as the reading on the nano-balance (in u). Use the **Exponent** slider to help get the correct amount. Stop adding atoms when the readings on both balances match exactly (to the nearest 0.001 g).

How many atoms did you need to add? _____

6.0221×10^{23}

2. **Explore:** Repeat the same procedure with carbon, then sulfur and aluminum.

A. For each element, how many atoms did you need to add? _

6.0221×10^{23}

B. What do you notice about the number of atoms in one mole? _

The number of atoms is the same in each element per mole

3. **Discover:** In each case, you measured out one mole of atoms, since the mass of one mole of any element, in grams, is equal to its atomic mass, in u. One mole of any element contains the same number of atoms, a number known as the **Avogadro constant**.

What is the exact value of the Avogadro constant? _____

$6.02214076 \times 10^{23}$

4. **Illustrate:** The Avogadro constant is so large it is normally written in **scientific notation**. To get an idea of the enormity of the Avogadro constant, write it out in standard form. (You will need to move the decimal place to the right 23 times, so you will need to add a lot of zeros!)

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