

**ANSWER KEY-Chem 200-Final Exam sample**

|    |                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                          |
|----|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| 7  | <p>A buffer solution contains 0.45 M sodium acetate (<math>\text{CH}_3\text{COONa}</math>) and 0.36 M acetic acid (<math>\text{CH}_3\text{COOH}</math>), <math>\text{p}K_a = 4.75</math>. What is the pH of this solution?</p> <p>A. 5.21<br/>B. 4.65<br/><b>C. 4.85</b><br/>D. 3.60</p>                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                 |
| 8  | <p>Which of the following acids is the <b>STRONGEST</b>?</p> <p>A. <b>HF</b> (<math>\text{p}K_a = 3.14</math>)    B. <math>\text{HNO}_2</math> (<math>\text{p}K_a = 3.35</math>)<br/>C. <math>\text{HCN}</math> (<math>\text{p}K_a = 9.4</math>)    D. <math>\text{HCOOH}</math> (<math>\text{p}K_a = 3.74</math>)</p>                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                   |
| 10 | <p>The following acids are listed with their <math>\text{p}K_a</math> values. Arrange the conjugate bases in order of <b>INCREASING</b> basic strength:<br/> <u>Conjugate acids:</u> <math>\text{NH}_4^+</math>    <math>(\text{CH}_3)_3\text{NH}^+</math>    <math>\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_3^+</math>    <math>(\text{CH}_3\text{CH}_2)_3\text{NH}^+</math><br/> <u><math>\text{p}K_a</math>:</u>                    9.2                    9.8                    10.5                    10.7</p> <p><b>A. <math>\text{NH}_3 &lt; (\text{CH}_3)_3\text{N} &lt; \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 &lt; (\text{CH}_3\text{CH}_2)_3\text{N}</math></b><br/>         B. <math>\text{NH}_3 &lt; \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 &lt; (\text{CH}_3)_3\text{N} &lt; (\text{CH}_3\text{CH}_2)_3\text{N}</math><br/>         C. <math>(\text{CH}_3\text{CH}_2)_3\text{N} &lt; \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2 &lt; (\text{CH}_3)_3\text{N} &lt; \text{NH}_3</math><br/>         D. <math>(\text{CH}_3\text{CH}_2)_3\text{N} &lt; (\text{CH}_3)_3\text{N} &lt; \text{NH}_3 &lt; \text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2</math></p> |
| 11 | <p>In order to detect the equivalence point in a titration, the chosen colored indicator should have a <math>\text{p}K_a</math> equal to</p> <p>A. pH at equivalence<br/>B. <b>pH at equivalence <math>\pm 1</math></b><br/>C. pH at equivalence + 1<br/>D. pH at equivalence - 1</p>                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                    |
| 12 | <p>Identify the reactant that is a <b>BRÖNSTED-LOWRY ACID</b> in the following reaction:    <math>\text{KCN} + \text{HI} \leftrightarrow \text{HCN} + \text{KI}</math></p> <p>A. KCN<br/>B. <b>HI</b><br/>C. HCN<br/>D. KI</p>                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                           |
| 13 | <p>Which of the following statements is <b>incorrect</b> at <math>25^\circ\text{C}</math>?</p> <p>A. <math>\text{pH} = -\log[\text{H}^+]</math><br/>B. <math>\text{pOH} = -\log[\text{OH}^-]</math><br/>C. <b><math>\text{pH} + \text{pOH} = 16</math></b><br/>D. <math>\text{pH} + \text{pOH} = \text{p}K_w</math></p>                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                  |
| 14 | <p>What is <b>TRUE</b> about <b>PURE</b> water?</p> <p>A. <math>\text{pH} = 7</math> at all temperatures<br/>B. <b><math>[\text{H}^+] = [\text{OH}^-]</math> at all temperatures</b><br/>C. <math>K_w = 10^{-14}</math> all temperatures<br/>D. All of the above</p>                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                     |

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