

Dr. Aquino

Name: _____

Please answer all questions in the space provided. A periodic table and formulas are provided at the end of the exam. You have 50 minutes. Calculators are permitted.

[9] Question 1

Urea (NH_2CONH_2) is the end-product in protein metabolism in animals. The decomposition of urea in 0.1 M HCl occurs as follows:



The reaction is *first-order* in urea and *first-order* overall. When $[\text{NH}_2\text{CONH}_2] = 0.400 \text{ M}$, the rate at 61°C is $1.71 \times 10^{-4} \text{ M/s}$.

a) What is the value for the rate constant, k ? [2]

$$\begin{aligned} \text{rate} &= k [\text{urea}] & k &= \frac{\text{rate}}{[\text{urea}]} = \frac{1.71 \times 10^{-4}}{0.400} \\ & & &= 4.28 \times 10^{-4} \text{ s}^{-1} \end{aligned}$$

b) What is the concentration of urea after 2.00×10^3 seconds, if the starting concentration is 0.700 M ? [3]

1st order

$$\begin{aligned} \ln [\text{urea}] &= -kt + \ln [\text{urea}]_0 \\ &= -(4.28 \times 10^{-4} \text{ s}^{-1})(2.00 \times 10^3 \text{ s}) + \ln (0.700) \\ &= -0.856 + (-0.357) \\ &= -1.213 \end{aligned}$$

$$[\text{urea}] = e^{-1.213} = \boxed{0.297 \text{ M}}$$

c) What is the half-life for this reaction at 61°C ? [3]
not a factor

$$t_{1/2} = \frac{0.693}{k} = \frac{0.693}{4.28 \times 10^{-4} \text{ s}^{-1}} = \frac{1620 \text{ sec}}{(1.62 \times 10^3 \text{ sec})}$$

d) Would the half-life be *shorter* or *longer* at 85°C ? [1]

Shorter

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