STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

Speed of light = (frequency)(wavelength)

 $c = f\lambda$

Energy = (Planck's constant)(frequency)

 $E_{photon} = hf$

Energy = (Planck's constant)(speed of light)
(wavelength)

 $E_{photon} = \frac{hc}{\lambda}$

BEHAVIOR OF GASES

Total pressure of a gas = (sum of the partial pressures)

 $P_{\mathrm{T}} = P_{1} + P_{2} + P_{3} + \dots$

(Pressure)(volume) = (moles)(ideal gas constant)(temperature)

PV = nRT

(Initial pressure)(initial volume)
(Initial moles)(initial temperature) = (final pressure)(final volume)

 $\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$

(Initial pressure)(initial volume) = (final pressure)(final volume)

 $P_1V_1 = P_2V_2$

(Initial volume) = (final volume)
(Initial temperature)

 $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

(Initial volume) = (final volume) (final moles)

 $\frac{V_1}{n_1} = \frac{V_2}{n_2}$

SOLUTIONS

Molarity = moles of solute liter of solution $M = \frac{\text{mod}}{1}$

Ionization constant of water = $\binom{hydrogen ion}{concentration}$ $\binom{hydroxide ion}{concentration}$

 $K_{w} = [H^{+}][OH^{-}]$

Volume of (molarity of solution 1) (molarity of solution 2) (molarity of solution 2)

 $V_1 M_1 = V_2 M_2$

pH = -logarithm (hydrogen ion concentration)

 $pH = -log[H^+]$

THERMOCHEMISTRY

Heat gained or lost = (mass) (specific) (change in temperature)

 $Q = mc_o \Delta T$

Enthalpy of = (enthalpy of products) - (enthalpy of reactants)

 $\Delta H = \Delta H_I^0$ (products) - ΔH_I^0 (reactants)

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Paul Warren

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The Journal of Physical Chemistry: Volume 67, Number 1 January 22, 1963 W. Albert Noyes, Ir., 1963

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