

Use the standard enthalpy of formation (ΔH_f°) values in **Appendix 1** to calculate the enthalpy of reaction for the following equations. Balance the equations where necessary.

- 5.
- (a) $\text{CH}_4 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{l})$
 - (b) $\text{C}_2\text{H}_6 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{l})$
 - (c) $\text{C}_{12}\text{H}_{22}\text{O}_{11} (\text{s, sucrose}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 (\text{s, glucose}) + \text{C}_6\text{H}_{12}\text{O}_6 (\text{s, fructose})$
 - (d) $\text{CaCO}_3 (\text{s}) \rightarrow \text{CaO} (\text{s}) + \text{CO}_2 (\text{g})$
 - (e) $\text{HCl} (\text{aq}) + \text{NaOH} (\text{aq}) \rightarrow \text{NaCl} (\text{aq}) + \text{H}_2\text{O} (\text{l})$
 - (f) $\text{N}_2\text{O}_5 (\text{g}) \rightarrow \text{N}_2 (\text{g}) + \text{O}_2 (\text{g})$

Appendix 1: Standard Enthalpy of Formation, ΔH_f° (kJ/mol).

Substance	ΔH_f° (kJ/mol)
$\text{CH}_4 (\text{g})$	-74.6
$\text{C}_2\text{H}_6 (\text{g})$	-84.0
$\text{C}_6\text{H}_{12}\text{O}_6 (\text{s, glucose})$	-1271.0
$\text{C}_6\text{H}_{12}\text{O}_6 (\text{s, fructose})$	-1266.6
$\text{C}_{12}\text{H}_{22}\text{O}_{11} (\text{s, sucrose})$	-2222.1
$\text{CaCO}_3 (\text{s})$	-1207.6
$\text{CaO} (\text{s})$	-6349.0
$\text{CO}_2 (\text{g})$	-393.5

Substance	ΔH_f° (kJ/mol)
$\text{H}_2\text{O} (\text{g})$	-241.8
$\text{H}_2\text{O} (\text{l})$	-285.8
$\text{HCl} (\text{aq})$	-167.2
$\text{N}_2 (\text{g})$	0.0
$\text{N}_2\text{O}_5 (\text{g})$	+13.3
$\text{NaCl} (\text{aq})$	-407.3
$\text{NaOH} (\text{aq})$	-44.5
$\text{O}_2 (\text{g})$	0.0

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Table of Contents Standard Enthalpy Of Formation Practice Questions

1. Understanding the eBook Standard Enthalpy Of Formation Practice Questions
 - The Rise of Digital Reading Standard Enthalpy Of Formation Practice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Standard Enthalpy Of Formation Practice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Standard Enthalpy Of Formation Practice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Standard Enthalpy Of Formation Practice Questions
 - Personalized Recommendations
 - Standard Enthalpy Of Formation Practice Questions User Reviews and Ratings
 - Standard Enthalpy Of Formation Practice Questions and Bestseller Lists
5. Accessing Standard Enthalpy Of Formation Practice Questions Free and Paid eBooks
 - Standard Enthalpy Of Formation Practice Questions Public Domain eBooks
 - Standard Enthalpy Of Formation Practice Questions eBook Subscription Services
 - Standard Enthalpy Of Formation Practice Questions Budget-Friendly Options
6. Navigating Standard Enthalpy Of Formation Practice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Standard Enthalpy Of Formation Practice Questions Compatibility with Devices
 - Standard Enthalpy Of Formation Practice Questions Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Standard Enthalpy Of Formation Practice Questions
 - Highlighting and Note-Taking Standard Enthalpy Of Formation Practice Questions
 - Interactive Elements Standard Enthalpy Of Formation Practice Questions
8. Staying Engaged with Standard Enthalpy Of Formation Practice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Standard Enthalpy Of Formation Practice Questions
 9. Balancing eBooks and Physical Books Standard Enthalpy Of Formation Practice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Standard Enthalpy Of Formation Practice Questions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Standard Enthalpy Of Formation Practice Questions
 - Setting Reading Goals Standard Enthalpy Of Formation Practice Questions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Standard Enthalpy Of Formation Practice Questions
 - Fact-Checking eBook Content of Standard Enthalpy Of Formation Practice Questions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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