

The Synthesis and Analysis of Aspirin

Acetylsalicylic acid is the chemical name for aspirin, the ubiquitous pain reliever. One of the compounds used in the synthesis of aspirin is salicylic acid, which is itself a pain reliever that was known to many ancient cultures, including the Native Americans who extracted it from willow tree bark. Salicylic acid is extremely bitter tasting, and frequent use can cause severe stomach irritation. The search for a milder form of this pain reliever led to the successful synthesis of acetylsalicylic acid by the German chemist Felix Hoffmann in 1893.

Your two primary objectives in this experiment will be to synthesize and analyze aspirin. There is more than one way to synthesize aspirin; in this experiment, you will react acetic anhydride with salicylic acid in the presence of phosphoric acid (which acts as a catalyst). The reaction equation is shown below.



You will conduct two tests of your synthesis to verify that you did indeed make aspirin, and to determine its relative purity. First, you will measure the melting temperature of a sample of your product, using a Melt Station. Second, you will use a Spectrophotometer to test the absorbance of salicylic acid impurity in your aspirin sample after it has been complexed with Fe^{3+} from iron (III) nitrate solution.

OBJECTIVES

In this experiment, you will

- Synthesize a sample of acetylsalicylic acid (aspirin).
- Calculate the percent yield of your synthesis.
- Measure the melting temperature of your aspirin sample.
- Conduct a spectrophotometric analysis of your aspirin sample.

Synthesis Of Aspirin Pre Lab Answers

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