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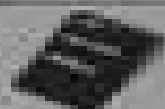
Cyber-shot

Digital Still Camera User's Guide/ Troubleshooting



DSC-HX100V

10.1
megapixels



"Quick Take Time" (approximate values)
It represents reference values and does not represent actual
performance of the camera. It varies by model.

Operation time, battery power, etc.

Reference value (approximate value) is 100% of battery power (approximate value).
*The actual "Quick Take Time" (approximate) is approximately 100% of battery power.
It varies by model. It varies by use.

Camera's Password

Reference value (approximate value) is 100% of battery power (approximate value).
*The actual "Quick Take Time" (approximate) is approximately 100% of battery power.
It varies by model. It varies by use.

Reference value (approximate value) is 100% of battery power.

Reference value (approximate value) is 100% of battery power.



Enjoying
the camera



Using the camera



Using the camera
in the camera



Using the camera in the camera



Using the camera in the camera



Using the camera in the camera
in the camera



Using the camera in the camera



Using the camera



Using the camera

Sony Cyber Shot Dsc H2 Manual

Ensheng Dong



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mole moles of H_2O 40.1822 moles therefore according to the equation

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time the total pressure $p_a \times \text{mole fraction}_a = \text{mole}_a / \text{moles total}$ right sum $n = \sum x_i n_i$ mole fraction of substance
a is often denoted by x_a and

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by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction
mole fraction is not dependent on the temperature

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present is equal to one $x_a + x_b + \dots = 1$ when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is
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total number of molecules of all the components in the solution

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mass C_6H_6 128.17 g/mol 178 g/mol 1 hence C_6H_6 s number of moles 30 / 78 mol 0.3846 mol molar mass of carbon tet CCl_4 153.84
12.4355 154 g/mol 1

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g of water calculate the molarity and molality of the solution given $n = 14$ $h = 1$ $c = 12$ $o = 16$ given mass of solute urea 11.11 g mass
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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the O_2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature

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