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ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ $ph = -\log 0.100 = 1.000$ b titrant volume 12.50 ml

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Volume of NaOH added (mL)	pH of the solution
4.6	4.6
4.6	4.6
5.9	5.9
5.9	5.9

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