

Target Stoichiometry Lab

Mole Relationships and the Balanced Equation



Introduction

A simple decomposition reaction of sodium bicarbonate (baking soda) presents the opportunity for students to test their knowledge of stoichiometry, limiting reagents, and the mole concept. This outcome-based lab requires the students to precisely predict the mass of the solid product. An electronic balance will reveal how close the students come to their "target" and determines their grades.

Concepts

- Mole concept
- Stoichiometry
- Balanced chemical equation

Materials

For each student group

- Sodium bicarbonate, NaHCO_3 , 5–6 g
- Universal indicator solution, a few drops
- Bunsen burner
- Butane safety lighter
- Clamps, single bar

To be shared

- Balances, electronic, 0.01-g precision, 3
- Test tube "holder" for balances (See Preparation below)
- For teacher preparation
- Glass wool

- Glass wool "paint brush" (See Preparation below)
- Support stand
- Test tube, borosilicate glass
- Test tube clamp

- Cardboard or thin metal sheets, 4" x 4", 3
- Drinking straw
- Electrical tape
- Scissors

Safety Precautions

Universal indicator solution is alcohol-based and is a flammable liquid. When heating sodium bicarbonate, use a borosilicate glass test tube. Inspect the test tube for chips and cracks before use and handle the hot test tube using a test tube clamp. Use a cool flame to heat the test tube. Do not handle glass wool with bare hands. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron. Wash hands thoroughly with soap and water before leaving the laboratory. Follow all laboratory safety guidelines. Please review current Material Safety Data Sheets for additional safety, handling, and disposal information.

Preparation

1. Make a glass wool "paint brush" for each student lab group. Wearing gloves, cut small pieces of glass wool and tape a piece to the end of each straw.
2. Fold a piece of cardboard or thin metal sheet accordion-style (see Figure 1). Place one folded sheet on each balance as a test tube holder to prevent the tube from rolling off the balance.
3. Set up a support stand and clamp with a Bunsen burner for each lab station.



Figure 1.

Disposal

Please consult your current *Flinn Scientific Catalog/Reference Manual* for general guidelines and specific procedures governing the disposal of laboratory waste. After completely cooling, the test tube of sodium carbonate may be rinsed and the waste disposed of down the drain with excess water according to Flinn Suggested Disposal Method #26b.

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Brendan G. Carr

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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