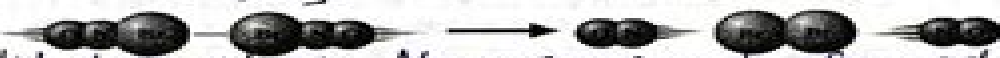


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1. Please answer the following questions in the space provided. [K5]

a) When solid KBr is dissolved in water, the solution gets colder. Is the chemical system exothermic or endothermic? [1]	endothermic
b) When a sample of potassium nitrate is dissolved in water, the temperature of the water changes from 23.5°C to 20.3°C. Is the chemical system exothermic or endothermic? [1]	endothermic
c) In the reaction $2A + B \rightarrow C + 3D$, reactant A is found to disappear at the rate of $6.2 \times 10^{-4} \text{ mol/L.s}$. What is the rate of disappearance of B? [1]	<u>2 mol</u> of A disappear at rate $6.2 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$ <u>1 mol</u> of B disappears at $x \frac{\text{mol}}{\text{L.s}}$ $x = \frac{6.2 \times 10^{-4} \text{ L.s}}{2} = 3.1 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$
d) The following diagram shows a collision between two BrNO molecules. Use collision theory to explain the results of the collision. [2]	 Reaction did take place. New compounds formed; 2 NO and 1 Br₂ are the products. It was an effective collision. Reactants collided with proper orientation and sufficient (optimal) kinetic energy.

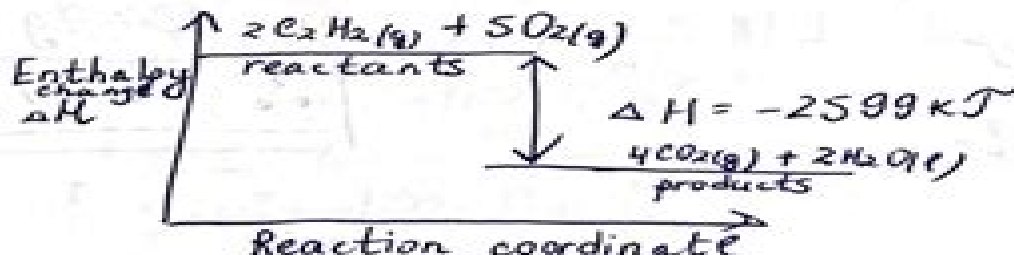
2. Consider the following reaction:



a) Write a thermochemical equation with the enthalpy change, ΔH , 'inside' the equation [C1]



b) Sketch an enthalpy diagram for this reaction. Provide appropriate labels for each axis [C3]



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Charles E. Mortimer



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Chemistry Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, 2004-07-01 Prentice Hall Chemistry meets the needs of students with a range of abilities diversities and learning styles by providing real world connections to chemical concepts and processes The first nine chapters introduce students to

the conceptual nature of chemistry before they encounter the more rigorous mathematical models and concepts in later chapters. The technology backbone of the program is the widely praised Interactive Textbook with ChemASAP which provides frequent opportunities to practice and reinforce key concepts with Animations Simulations Assessment and Problem solving tutorials

Sterling Test Prep MCAT General Chemistry Review: Complete Subject Review Sterling Test Prep, 2017-05-21 MCAT best seller used by thousands of students Higher score money back guarantee MCAT General Chemistry Complete Subject Review provides a detailed and thorough review of General Chemistry topics tested on the MCAT 2018. The content covers foundational principles and theories necessary to answer related questions on the test: Electronic structure and atomic nucleus, Periodic table, Chemical bonding, Stoichiometry, States of matter, Solution chemistry, Acids base equilibria, Rate processes in chemical reactions, Thermochemistry, thermodynamics, bioenergetics, Kinetics and equilibrium, Electrochemistry. This book provides a detailed and thorough review of General Chemistry topics tested on the MCAT. The content covers foundational principles and theories necessary to answer related questions on the test. The information is presented clearly and organized in a systematic way to provide students with a targeted MCAT review tool. You can focus on one knowledge area at a time to learn and fully comprehend important concepts and theories or to simply refresh your memory. By reading these review chapters thoroughly you will learn important chemistry concepts and the relationships between them so you can answer related questions on the test. This will prepare you for the MCAT general chemistry and you will significantly increase your score. Scoring well on the MCAT is extremely important for admission into medical school. To achieve a high MCAT score you need to develop skills to properly apply the knowledge you have and quickly choose the correct answer. Understanding key concepts, having the ability to extract information from the passages and distinguishing between similar answer choices is more valuable than simply memorizing terms. All the material in this book is prepared by science instructors with years of experience in applied chemistry and in academic settings. It was reviewed and organized by our MCAT editors to ensure strict adherence to the topics and skills outlined by the AAMC for the current MCAT. Our MCAT editors possess extensive credentials; they were educated in top colleges and universities and have been admitted to medical school with stellar MCAT scores. They are experts on teaching and preparing students for the MCAT and have coached thousands of premeds on admission strategies.

Answer Key and Test Manual Chemistry for the Health Sciences George I. Sackheim, Dennis D. Lehman, Ronald M. Schultz, 1977

A Level Chemistry MCQs Arshad Iqbal, 2017-04-20 A level chemistry multiple choice questions has 1749 MCQs. A level chemistry quiz questions and answers MCQs on A level chemistry: atomic structure, chemical bonding, chemistry of life, alcohols and esters, benzene, chemical compounds, analytical chemistry, MCQs with answers, carbonyl compounds, carboxylic acids, acyl compounds, electrode potential, electrons in atoms, enthalpy change, equilibrium, group IV, II and VII halogenoalkanes, hydrocarbon MCQs and quiz for SAT, ACT, GAT, GRE, CLEP, GED practice tests, AS level chemistry multiple choice quiz questions and answers, chemistry

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