Solubility Product, Ksp - Worksheet 1

1) Write the dissociation equation and the solubility product (Ksp) expression for the dissociation of each of the following substances.

a) PbSO₄ (s) b) Mg(OH)₂ (s) c) Pb(NO₃)₂ (s)d) AlCl₃ (s)

e) Na₂S (s) f) Ca₃(PO₄)₂(s)

Write the solubility product expression for the dissociation of each of the following solids in water.

a) sodium chloride

c) silver sufide

e) magnesium sulfite

b) calcium carbonate

d) aluminum hydroxide

f) barium phosphate

The solubility of NiS is 1.0 x 10⁻¹² mol/L. Calculate the value of K_{SP} for this compound.

4) The solubility of calcium hydroxide (Ca(OH)₂) is 0.02 mol/L. Calculate the solubility product constant for this substance.

5) The solubility of silver sulfide (Ag₂S) is 1.3 x 10⁻⁶ mol/L. Calculate the K_{sp} value.

6) Experiments show that 0.0059 g of SrCO₃ will dissolve in 1.0 L of water. What is the K_{8P} for strontium carbonate?

7) The solubility of cobalt (II) sulfide is 3.8 x 10⁻³ g/L. Calculate the value of Kap.

Solubility Product, K_{sp} - Worksheet 2

1. Calculate the solubility (in mol/L) for each of the following compounds:

a, FeS b, Mg(OH)₂ $K_{sp} = 3.7 \times 10^{-19}$ $K_{So} = 9.0 \times 10^{-12}$

c. Ag2S

 $K_{sp} = 1.6 \times 10^{-49}$

d. AgCN

 $K_{sp} = 2.0 \times 10^{-12}$

- 2. What is the solubility of silver iodide in grams per litre? $(K_{sp} = 8.3 \times 10^{-17})$
- The K_{sp} value for BaSO₄ is 1.5 x 10⁻⁹. Calculate the number of moles of this salt that will dissolve in 1000 mL of water.
- 4. Calculate the mass of CuS that will dissolve in 1.0 litre of water. $K_{sp} = 4.0 \times 10^{-38}$
- Calculate the mass of barium carbonate that will dissolve in 100 mL of water.
 The K_{8P} of BaCO₃ is 8.1 x 10⁻⁹.

Solubility Product Pogil Answers

Anna Maria Michałowska-Kaczmarczyk

Solubility Product Pogil Answers:

The Solubility Product Principle Sherry Lewin, 1960 **Solubility Equilibrium** Lifelige, 2019 This lesson plan covers the equilibrium of saturated solutions and how to calculate and use the solubility product constant **The Solubility Product Principle** Sherry Lewin, 1960 The Solubility Product Principle in Relation to the Dissociation of Salts Gilbert F. Smith.1961 The Thermodynamic Solubility Product and Its Importance in the Solubility Studies of Some **Quinidine Salts** Rajendra K. Khankari,1989 Solubility Products and Solubility Concepts Anna Maria Michałowska-Kaczmarczyk, 2017 The chapter refers to a general concept of solubility product Ksp of sparingly soluble hydroxides and different salts and calculation of solubility of some hydroxides oxides and different salts in aqueous media A criticized conventional approach based on stoichiometry of a reaction notation and the solubility product of a precipitate is compared with the unconventional correct approach based on charge and concentration balances and a detailed physicochemical knowledge on the system considered and calculations realized according to generalized approach to electrolytic systems GATES principles An indisputable advantage of the latter approach is proved in simulation of static or dynamic two phase nonredox or redox systems

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