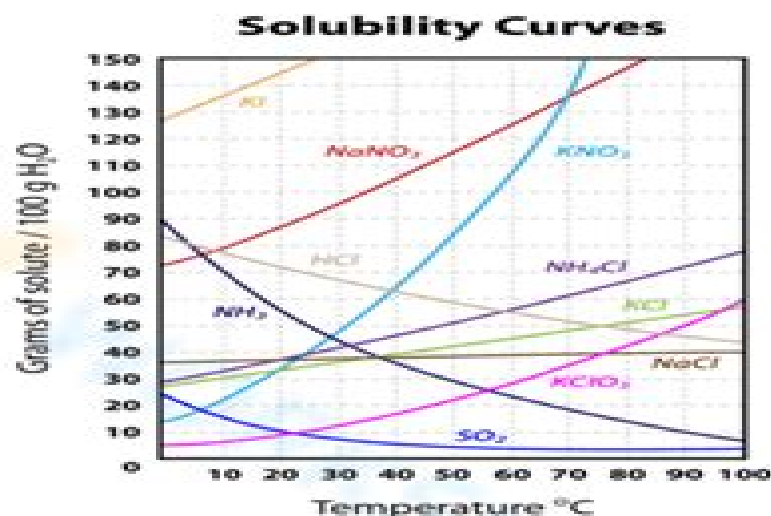


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 mL.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO₃ at 10°C, 40°C, & 80°C?
 @ 10°C = ~79 g
 @ 40°C = ~105 g
 @ 80°C = ~146 g
- How much NH₃ can you dissolve in water at 10°C, 30°C, & 90°C?
 @ 10°C = ~70 g
 @ 30°C = ~45 g
 @ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C, 50°C, & 90°C?
 @ 20°C = ~33 g
 @ 50°C = ~42 g
 @ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C?
About 40 grams
- Except for KI What substance is the most soluble in water at 20°C?
NaNO₃
- Except for KI, what substance is the most soluble in water at 72°C?
KNO₃
- Look up The boiling points for SO₂, HCl and NH₃ and write them below:
 B.P. SO₂ = -10°C
 B.P. HCl = -85°C
 B.P. NH₃ = -33°C
- Based on the boiling points you looked up, are SO₂, HCl & NH₃ solids, liquids, or gases at 0°C? (and anywhere else on the solubility curve)
gases

Solubility Curves Answers Chemistry If8766

William Henry Mickley



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Solubility curves Charles Hertz Frantz, 1933 **Solubility Curves** William Henry Mickley, 1933 **The Line**
Coordinate Representation of Solubility Curves and Reasons for Solubilities Dorothy Lee Pannell, 1946 **Principles**
of Solution and Solubility Kozo Shinoda, 1991 *Solubility Curves of Aqueous Or Aqueous Salt Solutions of Methyl*
Acetate-methanol and the Liquid-liquid Equilibrium of Methyl Acetate-methanol-aqueous Salt Solutions A. Nakamura, 1968

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Table of Contents Solubility Curves Answers Chemistry If8766

1. Understanding the eBook Solubility Curves Answers Chemistry If8766
 - The Rise of Digital Reading Solubility Curves Answers Chemistry If8766
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Curves Answers Chemistry If8766
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Curves Answers Chemistry If8766
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Curves Answers Chemistry If8766
 - Personalized Recommendations

- Solubility Curves Answers Chemistry If8766 User Reviews and Ratings
- Solubility Curves Answers Chemistry If8766 and Bestseller Lists
- 5. Accessing Solubility Curves Answers Chemistry If8766 Free and Paid eBooks
 - Solubility Curves Answers Chemistry If8766 Public Domain eBooks
 - Solubility Curves Answers Chemistry If8766 eBook Subscription Services
 - Solubility Curves Answers Chemistry If8766 Budget-Friendly Options
- 6. Navigating Solubility Curves Answers Chemistry If8766 eBook Formats
 - ePub, PDF, MOBI, and More
 - Solubility Curves Answers Chemistry If8766 Compatibility with Devices
 - Solubility Curves Answers Chemistry If8766 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Curves Answers Chemistry If8766
 - Highlighting and Note-Taking Solubility Curves Answers Chemistry If8766
 - Interactive Elements Solubility Curves Answers Chemistry If8766
- 8. Staying Engaged with Solubility Curves Answers Chemistry If8766
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Curves Answers Chemistry If8766
- 9. Balancing eBooks and Physical Books Solubility Curves Answers Chemistry If8766
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Curves Answers Chemistry If8766
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solubility Curves Answers Chemistry If8766
 - Setting Reading Goals Solubility Curves Answers Chemistry If8766
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solubility Curves Answers Chemistry If8766
 - Fact-Checking eBook Content of Solubility Curves Answers Chemistry If8766

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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