2.355a Cu (C2H3Oz)2.	L	mal Cu (Człły	g)2	1000 ML solu
2.355g Cn (C2H3Oz) 2 -	181.2	g Cu(CzH3¢	2)2	1 L sola
= 0.259504/	v = p	.260 M		
Sample Problem 2: Calculate the concentration in ppm. p				
copper (II) acetate (FW 181.5) and dil density of solution at the given tempe 4.70 to 104 ppm; 4.70 x 104 ppth		043 o/ml		flask. The
Sq Cu(Cathor)e Imbsolu	· · · · ·	g solut	- TX	)
· O art soln 1.0013 g sole	1.0	XIDE		
0. 6470388 g	Cu (Ortho)	$\omega = \frac{\times}{1.0 \times 0}$	٠ :	x=47,000
		1.0~1	13	47.03mppt

KEN

= 15.0 g Nacrt InvolNacrt # To.

= 15.0 g Nacrt InvolNacrt # To.

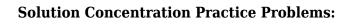
duriolized in approximately

0.75L, then

brought up to total Volume of 1.52

# **Solution Concentration Practice Problems**

**K Payea** 



#### Decoding **Solution Concentration Practice Problems**: Revealing the Captivating Potential of Verbal Expression

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