

Solubility Rules Chart

Negative Ions (Anions)	+	Positive Ions (Cations)	=	Solubility of Compounds in water	Example
any anion	+	Alkali Ions (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , Fr^+)	=	soluble	Sodium fluoride, NaF , is soluble
any anion	+	hydrogen ion [H^+ (aq)]	=	soluble	hydrogen chloride, HCl , is soluble
any anion	+	ammonium ion (NH_4^+)	=	soluble	ammonium chloride, NH_4Cl , is soluble
nitrate NO_3^-	+	any cation	=	soluble	potassium nitrate, KNO_3 , is soluble
acetate (CH_3COO^-)	+	any cation (except Ag^+)	=	soluble	sodium acetate, CH_3COONa , is soluble
Chloride (Cl^-), Bromide (Br^-), Iodide (I^-)	+	Ag^+ , Pb^{2+} , Hg^{2+} , Cu^+ , Ti^+	=	low solubility (insoluble)	silver chloride, AgCl , forms a white precipitate
	+	any other cation	=	soluble	potassium bromide, KBr , is soluble
Sulfate (SO_4^{2-})	+	Ca^{2+} , Sr^{2+} , Ba^{2+} , Ag^{2+} , Pb^{2+} , Ra^{2+} , Hg^{2+}	=	low solubility (insoluble)	barium sulfate, BaSO_4 , forms a white precipitate
	+	any other cation	=	soluble	copper sulfate, CuSO_4 , is soluble
sulfide (S^{2-})	+	alkali ions (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , Fr^+), alkali earth metals (Be^{2+} , Mg^{2+} , Ca^{2+} , Sr^{2+} , Ba^{2+} , Ra^{2+}), and H^+ (aq), and NH_4^+	=	soluble	magnesium sulfide, MgS , is soluble
	+	any other cation	=	low solubility (insoluble)	zinc sulfide, ZnS , is insoluble

Solubility Chart Chemistry 1

VM Jensen



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